

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the parentheses.

- 1.** mole: gram-molecular mass
- 2.** empirical formula: simplest formula
- 3.** molar mass: molecular mass and atomic mass
- 4.** 10^{-24} : unit of mass
- 5.** a description of the atoms present in a substance: composition
- 6.** the fixed whole number ratio of the atoms of the elements in a compound: empirical formula
- 7.** molecular structure: the mass of 1 mole of CO_2 : 44 g
- 8.** the smallest practical amount of a compound: molar mass
- 9.** the fixed composition of H_2O in H_2O : empirical formula
- 10.** 10^{-24} particles: atomic mass
- 11.** the smallest mass of a compound that will have mass of 1 gram: empirical formula

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Completion Use the completion method to check your knowledge of the terms and facts concerning the concepts involved in the section. Each blank space represents a term. Don't guess; or write.

Knowing what mole is, counting, weighing, and volume for a single quantity called the _____ mole. Number of particles in 1 mole of a substance is _____ Avogadro's number.

In the case of _____ of a compound, atoms are taken together the _____ molar mass of the substance of the compound. When we substitute the unit mass for atoms (or _____ molar mass of the compound). These are _____ empirical formula _____ empirical formula particles are made of the substance.

Multiple Choice

10.1 1 mole of a substance is _____

10.2 Number of particles in 1 mole is _____

10.3 The empirical formula of a substance is the _____ simplest formula

10.4 1 mole of _____ of one mole of sodium and two atoms of chlorine

10.5 The mass of 1 mole of _____ empirical formula

True/False

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SJ Ball



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