

# Writing Nuclear Equations

## Chem Worksheet 4-4

Name \_\_\_\_\_

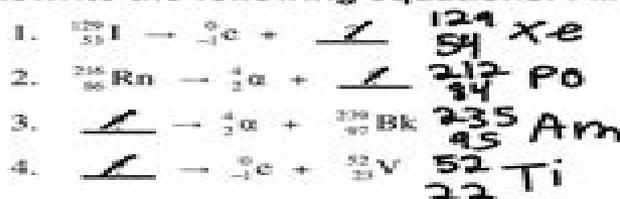
In the early 1900s scientists discovered that various isotopes will undergo nuclear decay. During this process the unstable nucleus of an atom gives off radiation. When scientists studied this radiation they discovered three types of particles: alpha, beta, and gamma. The **alpha particle** is composed of two protons and two neutrons, so it has a mass of 4 amu and a charge of 2+. A **beta particle** is a high energy electron emitted from the nucleus. A **gamma ray** often accompanies the other decay processes. Gamma radiation has no charge and no mass.

Radiation Type	Symbol	Mass (amu)	Charge
Alpha	${}^4_2\text{He}$ or ${}^4_2\alpha$	4	2+
Beta	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$	$\frac{1}{1840}$	1-
Gamma	${}^0_0\gamma$	0	0

Equations can be written to show how a nucleus changes during a nuclear decay process. With these nuclear equations we track the atomic number and the mass number. For this reason it is important to correctly write the symbols for each particle involved. A nuclear equation is written for an alpha decay and a beta decay below. Notice that the sum of the atomic numbers is equal on both sides of the arrow. The sum of the mass numbers is also the same on both sides.



Rewrite the following equations. Fill in all the missing information.



Write nuclear equations that describe the following processes.

- Uranium-235 undergoes an alpha decay to produce thorium-231.  $\rightarrow {}^{235}_{92}\text{U} \rightarrow {}^4_2\text{He} + {}^{231}_{90}\text{Th}$
- Lanthanum-144 becomes cerium-144 when it undergoes a beta decay.  $\rightarrow {}^{144}_{57}\text{La} \rightarrow {}^0_{-1}\text{e} + {}^{144}_{58}\text{Ce}$
- Neptunium-233 is formed when americium-237 undergoes a nuclear decay process.  $\rightarrow {}^{237}_{95}\text{Am} \rightarrow {}^{233}_{93}\text{Np} + {}^4_2\text{He}$
- When protactinium-229 goes through two alpha decays, francium-221 is formed.  $\rightarrow {}^{229}_{91}\text{Pa} \rightarrow {}^4_2\text{He} + {}^4_2\text{He} + {}^{221}_{87}\text{Fr}$
- Uranium-238 undergoes an alpha decay and produces two gamma rays.  $\rightarrow {}^{238}_{92}\text{U} \rightarrow {}^4_2\text{He} + {}^0_0\gamma + {}^0_0\gamma$
- The neon-22 nucleus is formed when an element undergoes a beta decay.  $\rightarrow {}^{22}_{11}\text{Na} \rightarrow {}^0_{-1}\text{e} + {}^{22}_{10}\text{Ne}$
- Samarium-146 is produced when an element undergoes an alpha decay.  $\rightarrow {}^{150}_{64}\text{Gd} \rightarrow {}^4_2\text{He} + {}^{146}_{62}\text{Sm}$
- The beta decay of dysprosium-165 creates a new element.  $\rightarrow {}^0_{-1}\text{e} + {}^{165}_{65}\text{Dy} \rightarrow {}^{165}_{66}\text{Tb}$

Answer the following questions. Include the mass number when naming isotopes.

- What atom produces scandium-47 when it goes through a beta decay?  $\rightarrow$  Titanium 47
- What new element is formed when curium-244 emits two alpha particles and three gamma rays?

Uranium  $\downarrow$   
236

# John Erickson 2005 Answer Key

**K Morrison**



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## **Table of Contents John Erickson 2005 Answer Key**

1. Understanding the eBook John Erickson 2005 Answer Key
  - The Rise of Digital Reading John Erickson 2005 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying John Erickson 2005 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an John Erickson 2005 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from John Erickson 2005 Answer Key
  - Personalized Recommendations
  - John Erickson 2005 Answer Key User Reviews and Ratings
  - John Erickson 2005 Answer Key and Bestseller Lists
5. Accessing John Erickson 2005 Answer Key Free and Paid eBooks
  - John Erickson 2005 Answer Key Public Domain eBooks
  - John Erickson 2005 Answer Key eBook Subscription Services
  - John Erickson 2005 Answer Key Budget-Friendly Options
6. Navigating John Erickson 2005 Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - John Erickson 2005 Answer Key Compatibility with Devices
  - John Erickson 2005 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of John Erickson 2005 Answer Key
  - Highlighting and Note-Taking John Erickson 2005 Answer Key
  - Interactive Elements John Erickson 2005 Answer Key
8. Staying Engaged with John Erickson 2005 Answer Key

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers John Erickson 2005 Answer Key
9. Balancing eBooks and Physical Books John Erickson 2005 Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection John Erickson 2005 Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine John Erickson 2005 Answer Key
- Setting Reading Goals John Erickson 2005 Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of John Erickson 2005 Answer Key
- Fact-Checking eBook Content of John Erickson 2005 Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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