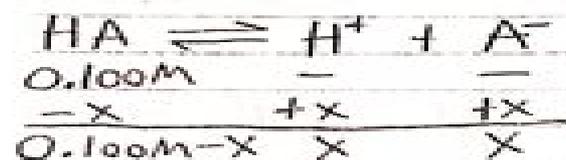


• Calculate  $K_a$  for the following acids:

1) 100.0 mL of 0.100 M HA with a pH of 3.40

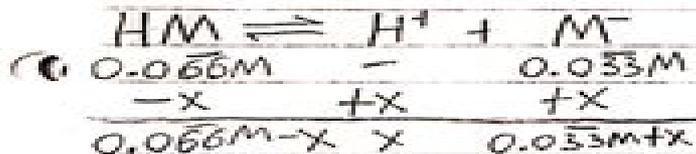


$$\text{pH} = 3.40 \rightarrow \text{pH} = -\log[\text{H}^+] \\
 3.40 = -\log(x) \rightarrow x = 3.98 \times 10^{-4}$$

$$K_a = \left( \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \right) = \left( \frac{(x)^2}{(0.100\text{M}-x)} \right)$$

$$K_a = \left( \frac{(3.98 \times 10^{-4})^2}{(0.100\text{M} - 3.98 \times 10^{-4})} \right) \rightarrow \boxed{K_a = 1.59 \times 10^{-6}}$$

2) 100.0 mL of 0.100 M HM is mixed with 50.0 mL of 0.100 M NaM and has a pH of 5.50



$$\text{pH} = 5.50 \rightarrow \text{pH} = -\log[\text{H}^+] \\
 5.50 = -\log(x) \rightarrow x = 3.16 \times 10^{-6}$$

Dilution problem due to HM and NaM

$$\begin{aligned}
 (C_1)(V_1) &= (C_2)(V_2) \\
 (0.100\text{M HM})(100.0\text{mL}) &= (C_2)(150.0\text{mL}) \\
 (C_2) &= 0.066\text{M HM}
 \end{aligned}$$

Dilution problem due to NaM and HM

$$\begin{aligned}
 (C_1)(V_1) &= (C_2)(V_2) \\
 (0.100\text{M NaM})(50.0\text{mL}) &= (C_2)(150.0\text{mL}) \\
 (C_2) &= 0.033\text{M NaM}
 \end{aligned}$$

$$K_a = \left( \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \right) = \left( \frac{(x)(0.033\text{M}+x)}{(0.066\text{M}-x)} \right)$$

$$K_a = \left( \frac{(3.16 \times 10^{-6})(0.033\text{M} + 3.16 \times 10^{-6})}{(0.066\text{M} - 3.16 \times 10^{-6})} \right) \rightarrow K_a = 1.58 \times 10^{-6}$$

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