

## Balancing Chemical Equations

Once we have written a chemical equation, we must make sure to balance it! Balancing a chemical equation means to make sure that there are the same number of each type of atom on both sides of the chemical equation. So, if there are five carbons on the left side (reactants), there needs to be five carbons on the right side (products).

So, how do we balance an equation? Follow the following steps...

- Change **only** stoichiometric coefficients – those big numbers in front of the compounds (not subscripts in the formulas – that would change the identity of the compound!)
- Start with most complex
- If found in more than one place on each side, do it last
- Watch out for odd numbers (make them into even ones if needed)
- Check your work!

Now, try to balance the following equations. In addition, name the **bolded** compounds for some nomenclature practice!

- $\underline{\quad} \text{NH}_4\text{NO}_3(\text{s}) \rightarrow \underline{\quad} \text{N}_2\text{O}(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{g})$   
*ammonium nitrate & dinitrogen monoxide*
- $\underline{\quad} \text{CH}_3\text{CO}_2\text{H}(\text{aq}) + \underline{2} \text{O}_2(\text{g}) \rightarrow \underline{2} \text{CO}_2(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{l})$   
*oxygen & carbon dioxide*
- $\underline{2} \text{Pb}(\text{NO}_3)_2(\text{s}) \rightarrow \underline{2} \text{N}_2\text{O}_4(\text{g}) + \underline{2} \text{PbO}(\text{s}) + \underline{\quad} \text{O}_2(\text{g})$   
*lead(II) nitrate & dinitrogen tetroxide & lead(II) oxide*
- $\underline{\quad} \text{SiO}_2(\text{s}) + \underline{4} \text{HF} \rightarrow \underline{\quad} \text{SiF}_4(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{l})$   
*silicon dioxide & silicon tetrafluoride*
- $\underline{8} \text{H}_2\text{S}(\text{aq}) + \underline{8} \text{Cl}_2(\text{g}) \rightarrow \underline{\quad} \text{S}_8(\text{s}) + \underline{16} \text{HCl}(\text{aq})$   
*chlorine*
- $\underline{\quad} \text{H}_2\text{CO}_3(\text{aq}) + \underline{\quad} \text{Ba}(\text{OH})_2 \rightarrow \underline{\quad} \text{BaCO}_3(\text{s}) + \underline{2} \text{H}_2\text{O}(\text{l})$   
*barium hydroxide & barium carbonate*
- $\underline{\quad} \text{P}_4(\text{s}) + \underline{6} \text{Br}_2(\text{l}) \rightarrow \underline{4} \text{PBr}_3(\text{l})$   
*phosphorous & phosphorous tribromide*
- $\underline{\quad} \text{S}_8(\text{l}) + \underline{4} \text{Cl}_2(\text{g}) \rightarrow \underline{4} \text{S}_2\text{Cl}_2(\text{g})$   
*disulfur dichloride*
- $\underline{2} \text{B}_5\text{H}_9(\text{l}) + \underline{12} \text{O}_2(\text{g}) \rightarrow \underline{5} \text{B}_2\text{O}_3(\text{s}) + \underline{9} \text{H}_2\text{O}(\text{g})$   
*pentaboron nonahydride & diboron trioxide*
- $\underline{3} \text{H}_2\text{SO}_4(\text{aq}) + \underline{2} \text{Al}(\text{OH})_3(\text{s}) \rightarrow \underline{\quad} \text{Al}_2(\text{SO}_4)_3(\text{aq}) + \underline{6} \text{H}_2\text{O}(\text{l})$   
*aluminum hydroxide & aluminum sulfate*
- $\underline{2} \text{Al}(\text{s}) + \underline{3} \text{Br}_2(\text{l}) \rightarrow \underline{2} \text{AlBr}_3(\text{s})$   
*aluminum bromide*
- $\underline{\quad} (\text{NH}_4)_2\text{Cr}_2\text{O}_7(\text{s}) \rightarrow \underline{\quad} \text{N}_2(\text{g}) + \underline{\quad} \text{Cr}_2\text{O}_3(\text{s}) + \underline{4} \text{H}_2\text{O}(\text{l})$   
*ammonium dichromate & chromium(III) oxide*
- $\underline{\quad} \text{S}_8(\text{s}) + \underline{24} \text{F}_2(\text{g}) \rightarrow \underline{8} \text{SF}_6(\text{g})$   
*sulfur hexafluoride*
- $\underline{\quad} \text{Zn}(\text{s}) + \underline{\quad} \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \underline{\quad} \text{Zn}(\text{NO}_3)_2(\text{aq}) + \underline{\quad} \text{Pb}(\text{s})$   
*zinc nitrate*

# Balancing Chemical Equations Multiple Choice Test

**Neil D. Jespersen, Pamela Kerrigan**



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