

✓✓✓✓✓  
experience

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or familiar with  
some experience

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suitable for competent  
DIY mechanics

✓✓✓  
for experienced DIY  
mechanics

✓✓✓  
suitable for expert  
DIY or professional

✓✓✓✓✓

## Specifications

### General

Manufacturer's engine codes:

up to model year 2012:

1800 cc 2.0 litre, 16-valve, DOHC

From approximately model year 2013 onward:

1800 cc 2.0 litre, 16-valve, DOHC

CAGA, CAGB, CAGC, CAHA, CAHB and CMEA

CGLC, CGLB, CACA, CABB, CACC, CACD, CMBA, CMFA,  
CMPB and CMDB

### Maximum outputs:

Engine code CAGA, CMEA, CACA and CMFA

Engine code CAGB and CACB

Engine code CAGC and CACC

Engine code CAHA

Engine code CAHB and CGLB

Engine code CGLC

Engine code CACD and CMPB

Engine code CMBA

### Power

100 kW @ 4200 rpm

100 kW @ 4200 rpm

88 kW @ 4200 rpm

129 kW @ 4200 rpm

120 kW @ 4200 rpm

130 kW @ 4200 rpm

110 kW @ 4200 rpm

130 kW @ 4200 rpm

### Torque

320 Nm @ 1750 to 2500 rpm

320 Nm @ 1750 to 2500 rpm

300 Nm @ 1750 to 2500 rpm

350 Nm @ 1750 to 2500 rpm

350 Nm @ 1750 to 2500 rpm

340 Nm @ 1750 to 2500 rpm

320 Nm @ 1750 to 2500 rpm

340 Nm @ 1750 to 2500 rpm

Bore

81.0 mm

Stroke

95.0 mm

Compression ratio:

Engine codes CAGA, CAGB, CAGC and CMEA

18.5 : 1

All other engines

16.5 : 1

Compression pressure:

Minimum compression pressure

Approximately 19.0 bar

Maximum difference between cylinders

Approximately 5.0 bar

Firing order

1 - 3 - 4 - 2

No 1 cylinder location

Timing belt end

**Note:** See 'Vehicle identification' in Reference chapter for the location of engine code markings.

### Lubrication system

Oil pump type

Gear type, driven by front balance shaft

Oil pressure switch (green cap)

0.5 bar

Oil pressure switch (brown cap)

0.7 bar

Oil pressure (oil temperature 80°C):

Minimum @ idling

0.8 bar

Minimum @ 2000 rpm

1.4 bar

# Audi Owners Manual

**JL Elias**



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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 and via of the periodic table sulfides can be formed by the direct reaction of the alkali metals with

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable

measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less

exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup> eg 2 HNO<sub>3</sub>(aq) + KOH(aq) → KNO<sub>3</sub>(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup>

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(aq) you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts