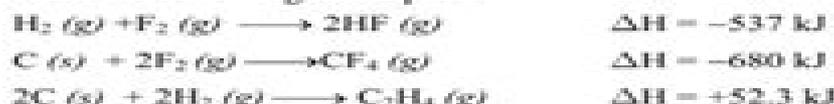


These problems are not meant to introduce the problems associated with thermochemistry. You should already have been introduced to the concepts in your lecture. These are just problems for extra practice. They are presented in random order, on purpose, to let you practice RECOGNIZING what TYPE of problem is in front of you, so that you know how to approach the problem.

Often you will find it much easier when you can identify the type of problem first.

1. From the following enthalpies of reaction:



calculate  $\Delta H$  for the reaction of ethylene with  $\text{F}_2$ :



2. When a 4.25-g sample of solid ammonium nitrate dissolves in 60.0 g of water in a coffee-cup calorimeter, the temperature drops from 22.0 °C to 16.9°C. Calculate  $\Delta H$  in (kJ/mol  $\text{NH}_4\text{NO}_3$ ) for the solution process  $\text{NH}_4\text{NO}_3 (s) \longrightarrow \text{NH}_4^+ (aq) + \text{NO}_3^- (aq)$ . Assume that the specific heat of the solution is the same as that of pure water.

3.  $\text{CH}_3\text{OH} (g) \longrightarrow \text{CO} (g) + 2\text{H}_2 (g) \quad \Delta H = +90.7 \text{ kJ}$

Calculate the amount of heat transferred when 45.0 g of  $\text{CH}_3\text{OH} (g)$  are decomposed by this reaction.

4. Calculate the standard enthalpy change for the following reactions:



substance	$\Delta H_f^\circ$ kJ/mol
$\text{SiCl}_4 (l)$	-640.1
$\text{SiO}_2 (s)$	-910.9
$\text{HCl} (aq)$	-167.2
$\text{HCl} (g)$	-92.30
$\text{H}_2\text{O} (g)$	-241.82
$\text{H}_2\text{O} (l)$	-285.83

5.  $2\text{Mg} (s) + \text{O}_2 (g) \longrightarrow 2\text{MgO} (s) \quad \Delta H = -1204 \text{ kJ}$

How many kJ of heat are absorbed when 7.50 g of  $\text{MgO} (s)$  are decomposed into  $\text{Mg} (s)$  &  $\text{O}_2 (g)$ ?

6. Write a balanced equation that corresponds to  $\Delta H_f^\circ$  for  $\text{SO}_2 (g)$ .

7. A 440. g quantity of a certain metal is heated to 100.0 °C. It is immediately thrust into 258 g of water that is initially at 25.0°C. The temperature of the metal-water mixture rises to 36.5 °C. What is the specific heat of the metal? Assume the density of the water is 0.998 g/mL and the specific heat capacity of the water is 4.184 J/g·°C.

# Thermochemistry Practice Test A Answers

**Peter J. Mikulecky, Michelle Rose  
Gilman, Kate Brutlag**



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## **Table of Contents Thermochemistry Practice Test A Answers**

1. Understanding the eBook Thermochemistry Practice Test A Answers
  - The Rise of Digital Reading Thermochemistry Practice Test A Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Practice Test A Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Thermochemistry Practice Test A Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Practice Test A Answers
  - Personalized Recommendations
  - Thermochemistry Practice Test A Answers User Reviews and Ratings
  - Thermochemistry Practice Test A Answers and Bestseller Lists
5. Accessing Thermochemistry Practice Test A Answers Free and Paid eBooks

- Thermochemistry Practice Test A Answers Public Domain eBooks
  - Thermochemistry Practice Test A Answers eBook Subscription Services
  - Thermochemistry Practice Test A Answers Budget-Friendly Options
6. Navigating Thermochemistry Practice Test A Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Thermochemistry Practice Test A Answers Compatibility with Devices
    - Thermochemistry Practice Test A Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Thermochemistry Practice Test A Answers
    - Highlighting and Note-Taking Thermochemistry Practice Test A Answers
    - Interactive Elements Thermochemistry Practice Test A Answers
  8. Staying Engaged with Thermochemistry Practice Test A Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Thermochemistry Practice Test A Answers
  9. Balancing eBooks and Physical Books Thermochemistry Practice Test A Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Thermochemistry Practice Test A Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Thermochemistry Practice Test A Answers
    - Setting Reading Goals Thermochemistry Practice Test A Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Thermochemistry Practice Test A Answers
    - Fact-Checking eBook Content of Thermochemistry Practice Test A Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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