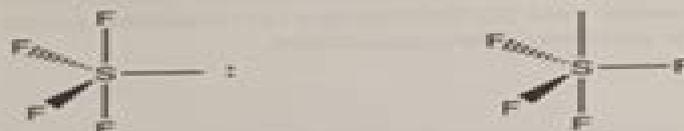


## Post Laboratory Assignment

Answer the following on a separate sheet of paper.

- Using the data collected for  $\text{PF}_5$ ,  $\text{SF}_6$  and  $\text{ClF}_3$ , what structural comments can be made about molecules with more than four groups (charge clouds) around the central atom?
- Select which structure is preferred and discuss why?



- To assess the stability of the axial vs. equatorial positions for a trigonal bipyramidal structure, a computer program was used to determine the energy for  $\text{SF}_4$  with the lone pair in either position (axial or equatorial). The data obtained is shown below:

Equatorial: Energy = -795.14 au

Axial: Energy = -795.088 au

- Using the conversion factor  $627.5096 \text{ kcal} = 1 \text{ au}$ , what is the energy difference in kcal between the two structures?
  - According to these data, which structure is more stable (more negative)?
  - Does this agree with your explanation from the previous question? Why or why not?
- Below is a table that lists dipole moments and bond angles calculated by a computer program for three molecules. Discuss why the bond angles differ and describe the significance for the dipole moment difference.

Molecule	Dipole moment	Bond Angle
Methane	0.00 Debye	$109.47^\circ$
Ammonia	1.84 Debye	$107.57^\circ$
Water	2.15 Debye	$105.93^\circ$

- Below is a table that lists dipole moments for three molecules. Discuss the ionic and covalent character of each compound using the dipole moment values.

Molecule	Dipole moment
$\text{H}_2$	0.00 Debye
HF	1.94 Debye
LiH	5.95 Debye

# Experiment 11 Molecular Models Answers

**Daniel F McAuley**



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