

# Covalent Bonding

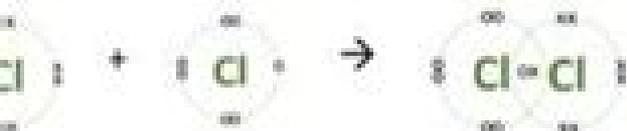
atoms need to get more electrons to have a full outer shell. They can do this in 2 ways:

**ionic bonding** - atoms lose or gain electrons, forming positive and negative ions which are then electrostatically attracted to each other (an ionic bond). This occurs between a metal and a non-metal.

**Covalent bonding** - atoms share a pair of electrons (a covalent bond). This occurs between non-metals only.

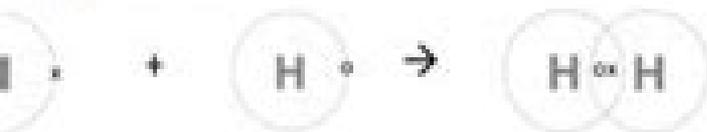
Bonding can only happen between non-metals - it does not involve metals.

atoms 'share' a pair of electrons. They do this by overlapping their outer shells so that the shared electrons are present in the shells of both atoms. This shared pair of electrons is a covalent bond.



note that as both atoms now possess the shared pair, both atoms have full outer shells of 8 electrons in their outer shell (in the covalent bond, 8 each on their outer shells).

hydrogen atoms are involved in covalent bonding, they will only have 2 electrons in their outer shell as the first shell can only contain 2 electrons. (remember their outer shell is still full).



## How covalent dot cross diagrams:

1. How many electrons do the atoms need to gain to have a full outer shell? This is how many bonds (x) go in the centre.

2. How many electrons does the atom have? If it has 7 electrons in its outer shell, it needs to gain 1 electron so it will form 1 covalent bond. This is how many dots (x) go in the centre.

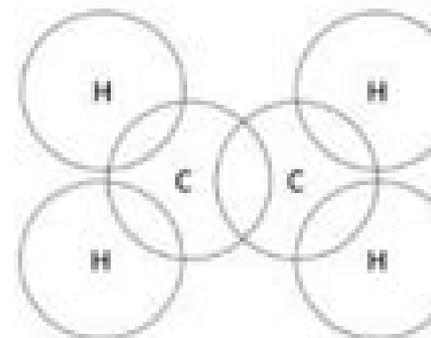
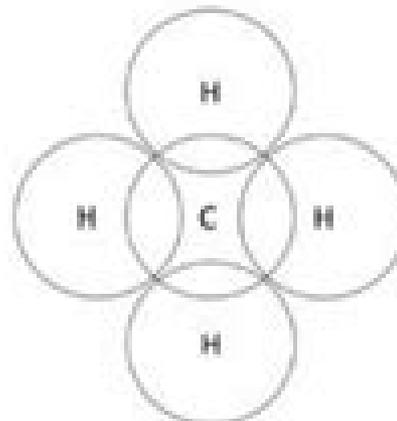
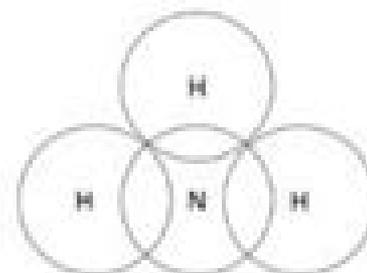
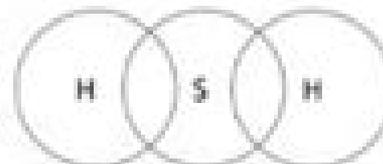
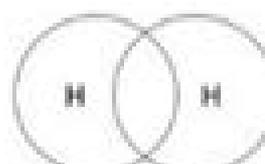
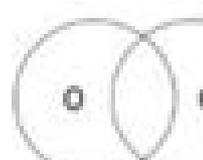
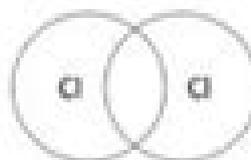


3. Add more electrons to the outer shell of each atom, so that they both have a full outer shell (in this case, 8 electrons). Remember that both atoms already 'own' the 2 shared electrons in the centre.

4. Each atom already has 2 electrons (from the shared pair), so they each need 6 more. Their outer shells are now full.



# Dot Cross Diagram Practice



# Dot Point Chemistry Answers Bonding

**John T. Moore, Richard H. Langley**



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