

Physical Chemistry Formula Revision

Basic Concepts

$$\text{Average atomic mass} = \frac{a_1x_1 + a_2x_2 + a_3x_3 + \dots}{100}$$

a_1, a_2, a_3, \dots - atomic mass of isotopes

x_1, x_2, x_3, \dots - percentage abundance

Mole concept

$$\text{no. of moles, } n = \frac{\text{given mass}}{\text{molar mass}} = \frac{W}{M}$$

$$n = \frac{\text{no. of particles}}{\text{Avogadro's No.}} = \frac{N}{6.022 \times 10^{23}}$$

$$n = \frac{\text{vol. of gas}}{\text{vol. of gas at STP}} = \frac{V}{22.4 \text{ L}}$$

Concentration Terms

$$\text{Percentage, } \% = \frac{W_1 \text{ or } V_1}{W \text{ or } V} \times 100$$

$$\text{Molarity, } M = \frac{n_1}{V(l)} = \frac{W_1 \times 1000}{M_1 \times V(\text{ml})}$$

$$\text{Normality, } N = \frac{n_1 \times 1000}{V(l)} = \frac{W_1 \times 1000}{E_1 \times V(\text{ml})}$$

$\hookrightarrow \frac{M_1}{n_f}$

$$\text{parts per million, ppm} = \frac{W_1}{W} \times 10^6$$

$$\text{Molality, } m = \frac{n_1}{W_2(\text{kg})} = \frac{W_1 \times 1000}{M_1 \times W_2(\text{g})}$$

$$\text{Mole fraction, } X_1 = \frac{n_1}{n_1 + n_2 + n_3 + \dots}$$

Stoichiometry



$$\frac{\text{moles of A}}{a} = \frac{\text{moles of B}}{b} = \frac{\text{moles of C}}{c}$$

Limiting R_1 : given amt. \rightarrow lowest stoichiometric coeff.

Atomic Structure

Planck's Quantum Theory

$$E = \frac{hc}{\lambda} \quad E(\text{eV}) = \frac{12400}{\lambda(\text{\AA})}$$

Photo Electric Effect



$$K_{\text{max}} = h\nu - \phi + K_{\text{emitted}}$$

$$e = h\nu_0 + \frac{m v_{\text{max}}^2}{2}$$



De Broglie's Hypothesis

$$m\lambda = \frac{h\nu}{\lambda} \rightarrow \lambda = \frac{h}{mv} \quad \lambda = \frac{h}{\sqrt{2mE}}$$

Bohr's Atomic model

$$mvr = \frac{nh}{2\pi} \quad r = 0.53 \times \frac{n^2}{Z} \text{ \AA}$$

$$r = 2.12 \times 10^{-8} \frac{n^2}{Z} \text{ m}$$

$$E_n = -13.6 \times \frac{Z^2}{n^2} \text{ eV} \quad T_n = \frac{h}{mv} = \frac{h}{2\pi mvr} = \frac{h}{2\pi E_n}$$

$$\frac{1}{\lambda} = R_H Z^2 \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right] \quad R_H = 1.097 \times 10^7 \text{ m}^{-1}$$

max no. of spectral lines when e^- returns from $n_2 = n_1 + 1$ $(2n_2 - n_1) \times (n_2 - n_1) / 2$

Heisenberg's uncertainty principle: $\Delta x \Delta p \geq \frac{h}{4\pi}$

Quantum Model

$$\text{total nodes} = n - 1$$

$$\text{angular nodes} = l$$

$$\text{radial nodes} = n - l - 1$$

Angular momentum

$$\text{orbital} = \sqrt{l(l+1)} \frac{h}{2\pi}$$

$$\text{spin} = \sqrt{s(s+1)} \frac{h}{2\pi}$$

Dictionary Of Physical Chemistry

Marco Cascella



Dictionary Of Physical Chemistry:

Reviewing **Dictionary Of Physical Chemistry**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Dictionary Of Physical Chemistry**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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