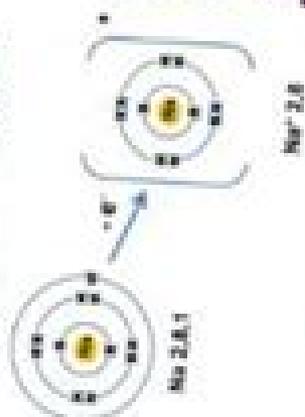


Group 1: The Alkali Metals

All have **one electron** in their **outermost shell**. React by **losing** this **electron** to form a **positive metal ion** with a **1+ charge**.



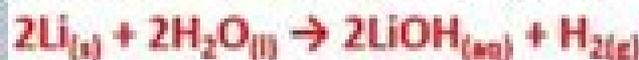
Reactivity increases as you go down the group. This is because:

- **Outermost electron** is in a shell which is **further from the nucleus** so it is **less strongly attracted** and **lost more easily**.
- **Increased shielding** as **more inner shell electrons** so is **less strongly attracted**.

3	Li	6.941
11	Na	22.990
19	K	39.100
37	Rb	85.467
55	Cs	132.9
87	Fr	223

LEAST REACTIVE
↓
MOST REACTIVE

All react with water to form an **alkaline** solution of **metal hydroxide** and **hydrogen gas**.



You will often be asked what you can **observe**:

- **The metal floats**
- **Fizzing / A gas is produced**
- **The metal 'disappears'**
- **The reactions become more vigorous as you do down the group**

As you do down group 1:

- **Melting point decreases**
 - They **become softer** / easier to cut.
- Increased distance** and **shielding** makes the **metallic bonding weaker**.



We call them alkali metals because they react with water to form **alkaline** solutions of **hydroxide ions (OH⁻)**. **Universal indicator** will turn **purple**. These solutions are highly **corrosive**



Alkali metals produce characteristic **colours in a flame**. We can use these to identify them:

- **Lithium – crimson**
- **Sodium – yellow**
- **Potassium - lilac**



Alkali metals react with oxygen to form **metal oxides**. Metal oxides are **bases** that react with **acids** and **neutralise** them.



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