

MATH WORKBOOK

Step by Step Guide
and Essential Practice
with Answers



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✓
Division

Place Value and Expanded Notations
✓

Fractions and Geometry
✓

Unit Conversion
✓

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5 ten millions + 1 hundred thousand
+ 3 ten thousands + 1 thousand
+ 1 hundred + 9 tens + 3 ones + 9 tenths

$$\begin{array}{r} 85.39 \\ \times 1.44 \\ \hline + 34156 \\ + 34156 \\ + 8539 \\ \hline = 1229616 \end{array}$$

$$\frac{41}{16} = 2 \frac{9}{16}$$

41 ÷ 16 = 2 with a remainder of 9

8 8, 16, 24, 32, 40, 48, 56
7 7, 14, 21, 28, 35, 42, 49, 56 , LCM = 56

$$0.5 = \frac{0.5}{1} = \frac{0.5 \times 6}{1 \times 6} = \frac{3}{6} = \frac{1}{2}$$

$$2 \frac{6}{10} \div 6 \frac{6}{7} = \frac{13 \times 7}{5 \times 48} = \frac{13 \times 7}{5 \times 48} = \frac{91}{240}$$

$$2 \times 10 = 6 \times 20 = \frac{20}{10} = \frac{13}{7} \quad 6 \times 6 = 48$$



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 $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \text{O}_2 \text{ (g)}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} \rightarrow \text{C}_2\text{H}_6 \text{ (g)}$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)} + \text{O}_2 \text{ (g)} \rightarrow \text{CO}_2 \text{ (g)} + \text{H}_2\text{O (g)}$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4 \text{ (g)} + 3 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 2 \text{ H}_2\text{O (l)}$ $\Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6 \text{ (g)} + 7/2 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (l)}$ $\Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2 \text{ (g)} + 1/2 \text{ O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$ $\Delta h = 285 \text{ kJ/mole}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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