



basic education

Department:
Basic Education
REPUBLIC OF SOUTH AFRICA

PHYSICAL SCIENCES

EXAMINATION GUIDELINES

GRADE 11

2025

These guidelines consist of 32 pages.

Grade 11 Physical Science Final Exam Paper

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2}\text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ}$

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 $\text{C}_2\text{N}_2(\text{g}) + 2 \text{ O}_2(\text{g}) \rightarrow \text{N}_2\text{O}_5(\text{g})$
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web hess s law states that if a reaction is the sum of two or more other reactions the Δh for the overall process must be the sum of the Δh values of the constituent reactions enthalpy change Δh is independent of the path that a reaction follows to move from reactants

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web this activity provides a demonstration of hess law using three reactions the solubility naoh in water the solubility naoh in hcl and the reaction of a solution of hcl and a solution of naoh

hess s law worksheet answers lozon - Jan 08 2023

web hess s law worksheet answers 1 calculate h for the reaction $c_2h_4(g) + h_2(g) \rightarrow c_2h_6(g)$ from the following data $c_2h_4(g) + 3/2 o_2(g) \rightarrow 2 co_2(g) + 2 h_2o(l) \quad \Delta h = 1411 \text{ kJ}$ $c_2h_6(g) + 3 1/2 o_2(g) \rightarrow 2 co_2(g) + 3 h_2o(l) \quad \Delta h = 1560 \text{ kJ}$

7a first law enthalpy calorimetry and hess s law worksheet - Oct 05 2022

web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

hess law lab science notes - Apr 30 2022

web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and Δh values to determine the heat of reaction for the reaction $c_2h_4(g) + h_2(g) \rightarrow c_2h_6(g)$

experiment 9 enthalpy of reaction hess s law - Jul 14 2023

web thus for reaction 9 2 in which 1 mole of h_3o^+ and 1 mole of oh^- combine to form 2 moles of h_2o $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{products} - \Delta h_{reactants}$ is positive if heat is absorbed that is if $\Delta h_{products} > \Delta h_{reactants}$

hess s law lab worksheet hess s law enthalpy of reaction studocu - Nov 06 2022

web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

hess s law lab report thermodynamics and calorimetry - Dec 07 2022

web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in

solution eventually allowing an individual to prove or disprove Hess's Law. Introduction: Thermodynamics is the study of energy exchanged between the system and the surroundings.

beyondlabz 5 Hess's Law docx 3 12 Hess's Law in this - Mar 10 2023

web 3 12 Hess's Law in this experiment you will measure the amount of heat released in these three related exothermic reactions:

1. $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$
2. $\text{NaOH (s)} + \text{H}^+ \text{(aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{H}_2\text{O (l)}$
- 3.