

Solutions for Gravimetric Analysis Exercises

1. The terms in a reaction quotient are actually dimensionless ratios of actual concentrations (or pressures) divided by standard concentrations (or pressures). The standard state for solutes is a 1 M solution and for gases it is a pressure of 1 bar (~ 1 atm), so these are the units used. For liquids and solids the standard state is the pure substance, so these ratios are unity. The solvent in a dilute solution is approximated as a pure substance.
2. The equilibrium constant for a reaction is dimensionless because it equals the reaction quotient at equilibrium, and, as discussed in Problem 1, this ratio involves only dimensionless terms.

3. a) Adding the two equations together gives



- b) At equilibrium $K_3 = [\text{AgCl}(\text{aq})]$, so $[\text{AgCl}(\text{aq})] = 3.6 \times 10^{-7}$ M.

4. Determine the solubility of $\text{La}(\text{IO}_3)_3$ using K_{sp} for $\text{La}(\text{IO}_3)_3$ and a table of initial and equilibrium concentration terms.

| | | | | | |
|-------|-------------------------------|----------------------|-----------------------------|-----|-----------------------------|
| | $\text{La}(\text{IO}_3)_3(s)$ | \rightleftharpoons | $\text{La}^{3+}(\text{aq})$ | $+$ | $3\text{IO}_3^-(\text{aq})$ |
| init | | | 0 | | 0.050 |
| equil | | | x | | $0.050 + 3x$ |

where x equals the increase in the $[\text{La}^{3+}]$ as a result of the dissolving process.

$$\text{at equil, } K_{sp} = 1.0 \times 10^{-11} = [\text{La}^{3+}][\text{IO}_3^-]^3 = (x)(0.050 + 3x)^3 = (x)(0.050)^3$$

\downarrow
 assumes $3x \ll 0.050$

$$[\text{La}^{3+}] = \frac{1.0 \times 10^{-11}}{(0.050)^3} = 8.0 \times 10^{-9} \text{ M} \quad (\text{assumption valid})$$

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John Anderson Wilkinson



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Table of Contents Solutions For Gravimetric Analysis Exercises

1. Understanding the eBook Solutions For Gravimetric Analysis Exercises
 - The Rise of Digital Reading Solutions For Gravimetric Analysis Exercises
 - Advantages of eBooks Over Traditional Books
2. Identifying Solutions For Gravimetric Analysis Exercises
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions For Gravimetric Analysis Exercises
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions For Gravimetric Analysis Exercises
 - Personalized Recommendations
 - Solutions For Gravimetric Analysis Exercises User Reviews and Ratings
 - Solutions For Gravimetric Analysis Exercises and Bestseller Lists

5. Accessing Solutions For Gravimetric Analysis Exercises Free and Paid eBooks
 - Solutions For Gravimetric Analysis Exercises Public Domain eBooks
 - Solutions For Gravimetric Analysis Exercises eBook Subscription Services
 - Solutions For Gravimetric Analysis Exercises Budget-Friendly Options
6. Navigating Solutions For Gravimetric Analysis Exercises eBook Formats
 - ePub, PDF, MOBI, and More
 - Solutions For Gravimetric Analysis Exercises Compatibility with Devices
 - Solutions For Gravimetric Analysis Exercises Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solutions For Gravimetric Analysis Exercises
 - Highlighting and Note-Taking Solutions For Gravimetric Analysis Exercises
 - Interactive Elements Solutions For Gravimetric Analysis Exercises
8. Staying Engaged with Solutions For Gravimetric Analysis Exercises
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions For Gravimetric Analysis Exercises
9. Balancing eBooks and Physical Books Solutions For Gravimetric Analysis Exercises
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions For Gravimetric Analysis Exercises
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solutions For Gravimetric Analysis Exercises
 - Setting Reading Goals Solutions For Gravimetric Analysis Exercises
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solutions For Gravimetric Analysis Exercises
 - Fact-Checking eBook Content of Solutions For Gravimetric Analysis Exercises
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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