CO_3^{2-} (M)	HCO_3^- (M)	pH of the mixture CO_3^{2-} and HCO_3^-	pH of the mixture with brine (0.15 M Mg^{2+})	XRD
1	0	11.2	9.8	MHC
0.75	0.25	10.3	7.8	Mg-calcite
0.5	0.5	9.8	6.6	Calcite
0.25	0.75	9.4	6.4	Calcite
0	1	8.2	5.4	Calcite

Whilst the potential Ca^{2+} separation stage utilising sodium bicarbonate is interesting, CO_2 loss associated with the said process must be considered. The reaction with bicarbonate changes the chemistry of the solution, releasing 1 mol of CO_2 per mol of precipitated calcium carbonate, because of the hydrogen ions released during the precipitation (Eq 10 and Eq 11):

$$Ca^{2+} + HCO_3^- = CaCO_3 + H^+ \qquad \text{Eq 10}$$

$$H^+ + HCO_3^- = H_2CO_3 = H_2O + CO_2 \uparrow \qquad \text{Eq 11}$$

Ph Of Calcium Carbonate Solution

Camilla Rothe

Ph Of Calcium Carbonate Solution:

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