

Solutions for Gravimetric Analysis Exercises

1. The terms in a reaction quotient are actually dimensionless ratios of actual concentrations (or pressures) divided by standard concentrations (or pressures). The standard state for solutes is a 1 M solution and for gases it is a pressure of 1 bar (~ 1 atm), so these are the units used. For liquids and solids the standard state is the pure substance, so these ratios are unity. The solvent in a dilute solution is approximated as a pure substance.
2. The equilibrium constant for a reaction is dimensionless because it equals the reaction quotient at equilibrium, and, as discussed in Problem 1, this ratio involves only dimensionless terms.

3. a) Adding the two equations together gives



- b) At equilibrium $K_3 = [\text{AgCl}(\text{aq})]$, so $[\text{AgCl}(\text{aq})] = 3.6 \times 10^{-7}$ M.

4. Determine the solubility of $\text{La}(\text{IO}_3)_3$ using K_{sp} for $\text{La}(\text{IO}_3)_3$ and a table of initial and equilibrium concentration terms.



where x equals the increase in the $[\text{La}^{3+}]$ as a result of the dissolving process.

$$\text{at equil, } K_{sp} = 1.0 \times 10^{-11} = [\text{La}^{3+}][\text{IO}_3^-]^3 = (x)(0.050 + 3x)^3 = (x)(0.050)^3$$

\downarrow
 assumes $3x \ll 0.050$

$$[\text{La}^{3+}] = \frac{1.0 \times 10^{-11}}{(0.050)^3} = 8.0 \times 10^{-9} \text{ M} \quad (\text{assumption valid})$$

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