

How do I prepare a phosphate buffer solution with a specific pH?

I am a science fair student at Shaker Heights Middle School. I need to make a phosphate buffer solution for my experiment. I am testing the effects of growth hormones on green plant life, and the buffer is to be absorbed by a sponge to aid the hormones in working. I would appreciate any assistance that you can give me. The solution needs to be a .001 M phosphate buffer with a pH of 6.1.

B. Peterson 1/13/98

Recipes for preparing phosphate buffer solutions can be found in many texts and handbooks. For example, the CRC Handbook of Chemistry and Physics recommends preparing a pH 6.10 buffer by adding 6.8 mL of 0.1 M NaOH to 50 mL of 0.1 M KH_2PO_4 . You can then dilute the buffer 1:100 for your needs.

Here is an abbreviated phosphate buffer table taken from Clark's *Determination of Hydrogen Ions*:

Clark and Lubs standard buffer solutions. Add the indicated number of mL of 0.2 M NaOH to 50 mL of 0.2 M KH_2PO_4 , and dilute to 200 mL. Taken from *Determination of Hydrogen Ions*, William and Wilkins, Baltimore, 1928.

| pH | mL NaOH | pH | mL NaOH |
|-----|---------|-----|---------|
| 5.8 | 3.72 | 7.0 | 29.63 |
| 6.0 | 5.70 | 7.2 | 35.00 |
| 6.2 | 8.60 | 7.4 | 39.50 |
| 6.4 | 12.60 | 7.6 | 42.80 |
| 6.6 | 17.80 | 7.8 | 45.20 |
| 6.8 | 23.65 | 8.0 | 46.80 |

Be sure you're using the correct phosphate. There are several potassium phosphates which can NOT be used interchangeably in buffer solutions! You want KH_2PO_4 .

| Formula | Systematic name | Other names |
|--------------------------|--------------------------------|---|
| KH_2PO_4 | potassium dihydrogen phosphate | acid potassium phosphate monobasic potassium phosphate monopotassium phosphate potassium biphosphate |
| K_2HPO_4 | potassium hydrogen phosphate | dipotassium phosphate dibasic potassium phosphate |
| K_3PO_4 | potassium phosphate | tripotassium phosphate tribasic potassium phosphate |

If you can't find such a recipe, follow this procedure:

1. **Use the Henderson-Hasselbalch equation** to compute the ratio of the base form to acid form at the desired pH. The Henderson-Hasselbalch equation is

Phosphate Buffer Solution Preparation

Terry C. Jones



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