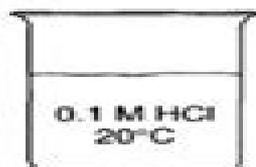
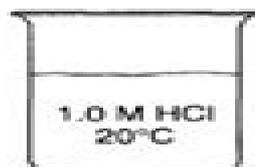


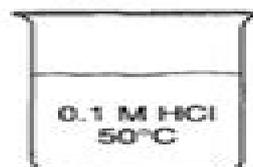
1. A reaction is most likely to occur when reactant particles collide with
  - A) proper energy, only
  - B) proper orientation, only
  - C) both proper energy and proper orientation**
  - D) neither proper energy nor proper orientation
2. The energy needed to start a chemical reaction is called
  - A) potential energy
  - B) kinetic energy
  - C) activation energy**
  - D) ionization energy
3. After being ignited in a Bunsen burner flame, a piece of magnesium ribbon burns brightly, giving off heat and light. In this situation, the Bunsen burner flame provides
  - A) ionization energy
  - B) activation energy**
  - C) heat of reaction
  - D) heat of vaporization
4. As the temperature increases, the rate of an exothermic reaction
  - A) decreases
  - B) increases**
  - C) remains the same
5. Which conditions will increase the rate of a chemical reaction?
  - A) decreased temperature and decreased concentration of reactants
  - B) decreased temperature and increased concentration of reactants
  - C) increased temperature and decreased concentration of reactants
  - D) increased temperature and increased concentration of reactants**
6. Base your answer to the following question on In each of the four beakers shown below, a 2.0-centimeter strip of magnesium ribbon reacts with 100 milliliters of  $\text{HCl}(\text{aq})$  under the conditions shown.



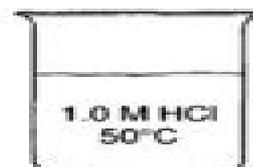
Beaker A



Beaker B



Beaker C



Beaker D

In which beaker will the reaction occur at the fastest rate?

- A) A
- B) B
- C) C
- D) D**

# Kinetics And Equilibrium Regents Questions Answers

**Rachel Sandford**



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