

Chapter Test B

Atomic Physics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. What is the frequency of a photon with an energy of $1.00 \times 10^{-19} \text{ J}$?
 ($h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$)
- | | |
|-------------------------------------|-------------------------------------|
| a. $1.00 \times 10^{14} \text{ Hz}$ | c. $3.00 \times 10^{12} \text{ Hz}$ |
| b. $2.00 \times 10^{14} \text{ Hz}$ | d. $4.00 \times 10^{12} \text{ Hz}$ |
- _____ 2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
- | | |
|----------|----------|
| a. 1 : 1 | c. 3 : 1 |
| b. 2 : 1 | d. 4 : 1 |
- _____ 3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
- | | |
|-----------|-----------|
| a. 4.6 eV | c. 1.4 eV |
| b. 4.8 eV | d. 2.4 eV |
- _____ 4. What causes the bright lines in the emission spectrum of an element to occur?
- Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
- _____ 5. What causes the dark lines in the absorption spectrum of an element to occur?
- Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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Table of Contents Holt Physics Chapter 17 Test

1. Understanding the eBook Holt Physics Chapter 17 Test
 - The Rise of Digital Reading Holt Physics Chapter 17 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Physics Chapter 17 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Physics Chapter 17 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Physics Chapter 17 Test
 - Personalized Recommendations
 - Holt Physics Chapter 17 Test User Reviews and Ratings
 - Holt Physics Chapter 17 Test and Bestseller Lists
5. Accessing Holt Physics Chapter 17 Test Free and Paid eBooks
 - Holt Physics Chapter 17 Test Public Domain eBooks
 - Holt Physics Chapter 17 Test eBook Subscription Services
 - Holt Physics Chapter 17 Test Budget-Friendly Options
6. Navigating Holt Physics Chapter 17 Test eBook Formats

- ePub, PDF, MOBI, and More
 - Holt Physics Chapter 17 Test Compatibility with Devices
 - Holt Physics Chapter 17 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Physics Chapter 17 Test
 - Highlighting and Note-Taking Holt Physics Chapter 17 Test
 - Interactive Elements Holt Physics Chapter 17 Test
 8. Staying Engaged with Holt Physics Chapter 17 Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Physics Chapter 17 Test
 9. Balancing eBooks and Physical Books Holt Physics Chapter 17 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Physics Chapter 17 Test
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Holt Physics Chapter 17 Test
 - Setting Reading Goals Holt Physics Chapter 17 Test
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Holt Physics Chapter 17 Test
 - Fact-Checking eBook Content of Holt Physics Chapter 17 Test
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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