

Physical Science Chapter 6 Test Review Sheet 2014-15

Assessments- This is the portion of the test which you, the student, will be permitted to retest on.

On the assessment, students will be required to:

1. Write out the names and formulas for a binary ionic compound.
2. Write out the names and formulas for a covalent (molecular) compound.

Students will know the following prefixes used for naming covalent/molecular compounds:

Mono-	1	Di-	2	Tri-	3
Tetra-	4	Penta-	5	Hexa-	6
Hepta-	7	Octa-	8	Nona-	9
Deca-	10				

Multiple Choice/Short Answer/Essay

1. Students will know, understand, and be able to apply the following terms:

electron dot diagram	ion	anion	cation
chemical bond	ionic bond	chemical formula	
covalent bond	molecule	polar covalent bond	metallic bond
polyatomic ion	semiconductor	molecular formula	empirical formula
hydrogen bond	formula unit	nonpolar covalent bond	

2. Students will know the composition, physical and chemical properties of ionic and covalent compounds.
3. Students will be able to predict and explain why elements lose or gain electrons to form ions.
4. Students will describe which groups in the periodic table are most willing to lose or gain electrons and explain why.
5. Students will know the different diatomic molecules as described in class (incomplete list in textbook)
6. Students will know what factors play a role in whether a molecular compound will be polar or nonpolar.

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Table of Contents Physical Science Chapter 6 Test

1. Understanding the eBook Physical Science Chapter 6 Test
 - The Rise of Digital Reading Physical Science Chapter 6 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 6 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 6 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 6 Test
 - Personalized Recommendations
 - Physical Science Chapter 6 Test User Reviews and Ratings
 - Physical Science Chapter 6 Test and Bestseller Lists
5. Accessing Physical Science Chapter 6 Test Free and Paid eBooks

- Physical Science Chapter 6 Test Public Domain eBooks
 - Physical Science Chapter 6 Test eBook Subscription Services
 - Physical Science Chapter 6 Test Budget-Friendly Options
6. Navigating Physical Science Chapter 6 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chapter 6 Test Compatibility with Devices
 - Physical Science Chapter 6 Test Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 6 Test
 - Highlighting and Note-Taking Physical Science Chapter 6 Test
 - Interactive Elements Physical Science Chapter 6 Test
 8. Staying Engaged with Physical Science Chapter 6 Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 6 Test
 9. Balancing eBooks and Physical Books Physical Science Chapter 6 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 6 Test
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physical Science Chapter 6 Test
 - Setting Reading Goals Physical Science Chapter 6 Test
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Science Chapter 6 Test
 - Fact-Checking eBook Content of Physical Science Chapter 6 Test
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

-
- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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