

# The Periodic Table and Periodic Law

## Section 6.1 Development of the Modern Periodic Table

In your textbook, read about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

octaves	atomic mass	atomic number	nine
elements	properties	Henry Moseley	eight
protons	periodic law	Dmitri Mendeleev	accepted

The table below was developed by John Newlands and is based on a relationship called the law of **(1)** \_\_\_\_\_. According to this law, the properties of the elements repeated every **(2)** \_\_\_\_\_ elements. Thus, for example, element two and element **(3)** \_\_\_\_\_ have similar properties. The law of octaves did not work for all the known elements and was not generally **(4)** \_\_\_\_\_.

1	2	3	4	5	6	7
H	Li	G	Bo	C	N	O
B	S	10	11	12	13	14
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to **(5)** \_\_\_\_\_. In his table, the elements were arranged according to increasing **(6)** \_\_\_\_\_. One important result of this table was that the existence and properties of undiscovered **(7)** \_\_\_\_\_ could be predicted.

The element in the modern periodic table are arranged according to increasing **(8)** \_\_\_\_\_, as a result of the work of **(9)** \_\_\_\_\_. This arrangement is based on number of **(10)** \_\_\_\_\_ in the nucleus of an atom of the element. The modern form of the periodic table results in the **(11)** \_\_\_\_\_, which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical **(12)** \_\_\_\_\_.

# Teaching Transparency The Periodic Table Answers

**AW Chickering**



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