

SOLUTIONS OF ACIDS AND BASES

- ① Write the chemical equation showing how each of the following dissolves in water.



- ② For each reaction type, complete the equations and balance them.

a. Making acids from nonmetal oxide

General form: nonmetal oxide + water \longrightarrow acid

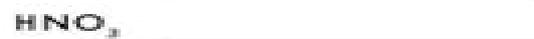


b. Making bases from metal oxide

General form: Metal oxide + water \longrightarrow base



- ③ Write the balanced equation for the titration between sulfurous acid and potassium hydroxide.
- ④ Write the balanced net ionic equation for the titration between sulfurous acid and potassium hydroxide.
- ⑤ Write an equation for the reaction of potassium metal with hydrochloric acid.
- ⑥ Acids can be labelled as monoprotic, diprotic, or triprotic. Label the following as monoprotic, diprotic, or triprotic



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Henry Verhoek**



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