

CHAPTER 7 PRACTICE TEST: ATOMIC STRUCTURE & PERIODIC TRENDS- KEY

1) When a solution of sodium chloride is vaporized in a flame, the color of the flame is:

- A) Blue **B) Yellow** C) Green D) Violet E) White

2) Electromagnetic radiation travels through vacuum at a speed of _____ m/s.

- A) 186,000 **C) 3.00×10^8** E) It depends on λ
 B) 125 D) 10,000 λ

3) The wavelength of light that has a frequency of $1.20 \times 10^{13} \text{ s}^{-1}$ is _____ m.

- A) 25.0 **B) 2.50×10^{-2}** C) 0.0400 E) 2.5
 D) 12.0

GIVEN: $\nu = 1.20 \times 10^{13} \text{ s}^{-1}$

UNKNOWN: λ

CONNECTION: $c = \lambda \nu$

SUBSTITUTE & SOLVE: $\lambda = c/\nu$

$$\lambda = \frac{3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}}{1.20 \times 10^{13} \text{ s}^{-1}}$$

$$\lambda = 2.50 \times 10^{-2} \text{ m}$$

4) Ham radio operators often broadcast on the 6-meter band. The frequency of this EM radiation is _____ MHz.

- A) 500 B) 200 **C) 50** D) 20 E) 2.0

GIVEN: $\lambda = 6 \text{ m}$

UNKNOWN: ν (Must convert answer to MHz!)

CONNECTION: $c = \lambda \nu$

SUBSTITUTE & SOLVE: $\nu = c/\lambda$

$$\nu = \frac{3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}}{6 \text{ m}}$$

$$\nu = 0.5 \times 10^8 \text{ Hz} \times \frac{1 \text{ MHz}}{10^6 \text{ Hz}} = 50 \text{ MHz}$$

5) The energy of a photon of light is _____ proportional to its frequency and _____ proportional to its wavelength.

- A) directly, directly C) inversely, directly E) indirectly, not
 B) inversely, inversely **D) directly, inversely**

6) The energy of a photon that has a wavelength of 12.3 nm is _____ J.

- A) 1.51×10^{-17} B) 4.42×10^{-23} C) 1.99×10^{-23} D) 2.72×10^{-20} **E) 1.62×10^{-17}**

7) The wavelength of a photon that has an energy of $5.25 \times 10^{-19} \text{ J}$ is _____ m.

- A) 3.79×10^{-7}** B) 2.64×10^6 C) 2.38×10^{23} D) 4.21×10^{-24} E) 3.79×10^7

8) A mole of red photons of wavelength 725 nm has _____ kJ of energy.

- A) 2.74×10^{-19} B) 4.56×10^{-26} C) 6.05×10^{-7} **D) 165** E) 227

GIVENS: $\lambda = 725 \text{ nm}$ (must convert to m before using formulas)

$$\lambda = 725 \text{ nm} \times \frac{1 \text{ m}}{10^9 \text{ nm}} = 7.25 \times 10^{-7} \text{ m}$$

Atomic Structure Practice Test Kent Chemistry

Michael Brown



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Table of Contents Atomic Structure Practice Test Kent Chemistry

1. Understanding the eBook Atomic Structure Practice Test Kent Chemistry
 - The Rise of Digital Reading Atomic Structure Practice Test Kent Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Practice Test Kent Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atomic Structure Practice Test Kent Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Practice Test Kent Chemistry

- Personalized Recommendations
 - Atomic Structure Practice Test Kent Chemistry User Reviews and Ratings
 - Atomic Structure Practice Test Kent Chemistry and Bestseller Lists
5. Accessing Atomic Structure Practice Test Kent Chemistry Free and Paid eBooks
 - Atomic Structure Practice Test Kent Chemistry Public Domain eBooks
 - Atomic Structure Practice Test Kent Chemistry eBook Subscription Services
 - Atomic Structure Practice Test Kent Chemistry Budget-Friendly Options
 6. Navigating Atomic Structure Practice Test Kent Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Atomic Structure Practice Test Kent Chemistry Compatibility with Devices
 - Atomic Structure Practice Test Kent Chemistry Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Atomic Structure Practice Test Kent Chemistry
 - Highlighting and Note-Taking Atomic Structure Practice Test Kent Chemistry
 - Interactive Elements Atomic Structure Practice Test Kent Chemistry
 8. Staying Engaged with Atomic Structure Practice Test Kent Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atomic Structure Practice Test Kent Chemistry
 9. Balancing eBooks and Physical Books Atomic Structure Practice Test Kent Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Atomic Structure Practice Test Kent Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Atomic Structure Practice Test Kent Chemistry
 - Setting Reading Goals Atomic Structure Practice Test Kent Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Atomic Structure Practice Test Kent Chemistry

- Fact-Checking eBook Content of Atomic Structure Practice Test Kent Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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