

Nuclear Chemistry Worksheet 3

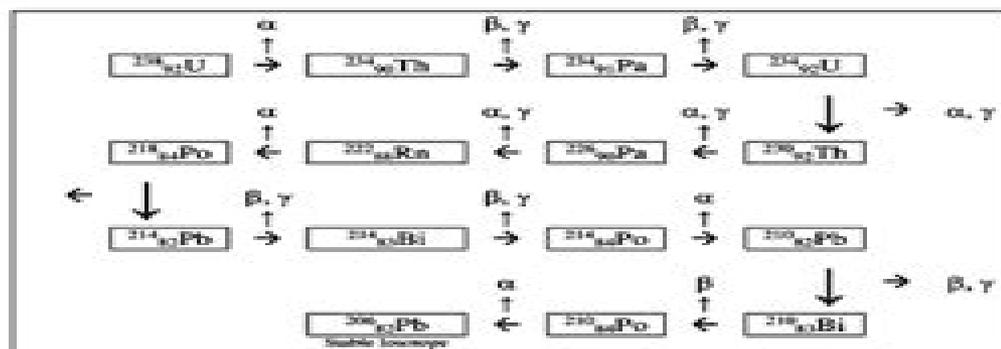
Directions: Identify the following as alpha, beta, gamma, or neutron.

- $\frac{1}{0}n$ _____
- $\frac{0}{-1}e$ _____
- $\frac{4}{2}He$ _____
- $\frac{0}{0}\gamma$ _____
- Nuclear decay with no mass and no charge _____
- An electron _____
- Least penetrating nuclear decay _____
- Most damaging nuclear decay to the human body _____
- Nuclear decay that can be stopped by skin or paper. _____
- Nuclear decay that can be stopped by aluminum. _____

Complete the following nuclear equations.

- ${}_{19}^{42}K \rightarrow {}_{-1}^0e + \underline{\hspace{2cm}}$
- ${}_{94}^{239}Pu \rightarrow \frac{4}{2}He + \underline{\hspace{2cm}}$
- ${}_{4}^9Be \rightarrow \frac{9}{4}Be + \underline{\hspace{2cm}}$
- ${}_{92}^{235}U \rightarrow \underline{\hspace{2cm}} + \frac{234}{90}Th$
- ${}_{3}^6Li \rightarrow \frac{4}{2}He + \underline{\hspace{2cm}}$
- $\underline{\hspace{2cm}} \rightarrow \frac{142}{56}Ba + \frac{91}{36}Kr + 3 \frac{1}{0}n$

Nuclear Decay Series



The figure maps the radioactive decay of uranium-238 to lead-206. Use the figure to answer the following questions.

- How many alpha particles are produced as one atom of uranium-238 decays to an atom of lead-206? _____
- How many beta particles? _____
- Explain why lead-206 is a stable isotope. _____

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Table of Contents 25 3 Nuclear Chemistry Review Answers Key

1. Understanding the eBook 25 3 Nuclear Chemistry Review Answers Key
 - The Rise of Digital Reading 25 3 Nuclear Chemistry Review Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 25 3 Nuclear Chemistry Review Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 25 3 Nuclear Chemistry Review Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 25 3 Nuclear Chemistry Review Answers Key
 - Personalized Recommendations
 - 25 3 Nuclear Chemistry Review Answers Key User Reviews and Ratings
 - 25 3 Nuclear Chemistry Review Answers Key and Bestseller Lists

5. Accessing 25 3 Nuclear Chemistry Review Answers Key Free and Paid eBooks
 - 25 3 Nuclear Chemistry Review Answers Key Public Domain eBooks
 - 25 3 Nuclear Chemistry Review Answers Key eBook Subscription Services
 - 25 3 Nuclear Chemistry Review Answers Key Budget-Friendly Options
6. Navigating 25 3 Nuclear Chemistry Review Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 25 3 Nuclear Chemistry Review Answers Key Compatibility with Devices
 - 25 3 Nuclear Chemistry Review Answers Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 25 3 Nuclear Chemistry Review Answers Key
 - Highlighting and Note-Taking 25 3 Nuclear Chemistry Review Answers Key
 - Interactive Elements 25 3 Nuclear Chemistry Review Answers Key
8. Staying Engaged with 25 3 Nuclear Chemistry Review Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 25 3 Nuclear Chemistry Review Answers Key
9. Balancing eBooks and Physical Books 25 3 Nuclear Chemistry Review Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 25 3 Nuclear Chemistry Review Answers Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 25 3 Nuclear Chemistry Review Answers Key
 - Setting Reading Goals 25 3 Nuclear Chemistry Review Answers Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 25 3 Nuclear Chemistry Review Answers Key
 - Fact-Checking eBook Content of 25 3 Nuclear Chemistry Review Answers Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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