

Mathematically

$$\frac{P^0 - P^s}{P^0} = x_2$$

Here  $P^0$  = Pressure of solvent in pure stage.

$P^s$  = Pressure of solution

$x_2$  = Mole fraction of solute.

### DERIVATION OF THE LAW

Pressure of solvent in solution

= Mole fraction of solvent  $\times$  pressure of solvent in pure stage

**If solute is non-volatile**

Total pressure of solution = Pressure due to solvent

= Mole fraction of solvent  $\times$  pressure of solvent in pure stage

So  $P_{\text{solvent}} = x_1 \cdot P_{\text{solution}}$  Here  $x_1$  is mole fraction of solvent.

$$\frac{P_{\text{solvent}}}{P_{\text{solution}}} = x_1$$

Subtract each term from 1 on both side.

$$1 - \frac{P_{\text{solvent}}}{P_{\text{solution}}} = 1 - x_1$$

$$\frac{P_{\text{solution}} - P_{\text{solvent}}}{P_{\text{solution}}} = x_2$$

### 2.11 IDEAL & NON-IDEAL SOLUTION

Sr.	Ideal solution	Non-Ideal solution
1.	A solution which obey Raoult's law is called ideal solution.	A solution which do not obey Raoult's law is called non-ideal solution.
2.	Interaction between components of mixture i.e.-between A&B is same as found in A-A & B-B molecule.	Interaction between components of mixture i.e.-between A&B is not same as found in A-A & B-B molecule.

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## **Table of Contents Class 12th Notes Chemistry Chapter Wise Notes**

1. Understanding the eBook Class 12th Notes Chemistry Chapter Wise Notes
  - The Rise of Digital Reading Class 12th Notes Chemistry Chapter Wise Notes
  - Advantages of eBooks Over Traditional Books
2. Identifying Class 12th Notes Chemistry Chapter Wise Notes
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Class 12th Notes Chemistry Chapter Wise Notes
  - User-Friendly Interface
4. Exploring eBook Recommendations from Class 12th Notes Chemistry Chapter Wise Notes
  - Personalized Recommendations
  - Class 12th Notes Chemistry Chapter Wise Notes User Reviews and Ratings
  - Class 12th Notes Chemistry Chapter Wise Notes and Bestseller Lists
5. Accessing Class 12th Notes Chemistry Chapter Wise Notes Free and Paid eBooks
  - Class 12th Notes Chemistry Chapter Wise Notes Public Domain eBooks
  - Class 12th Notes Chemistry Chapter Wise Notes eBook Subscription Services
  - Class 12th Notes Chemistry Chapter Wise Notes Budget-Friendly Options
6. Navigating Class 12th Notes Chemistry Chapter Wise Notes eBook Formats

- ePub, PDF, MOBI, and More
  - Class 12th Notes Chemistry Chapter Wise Notes Compatibility with Devices
  - Class 12th Notes Chemistry Chapter Wise Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Class 12th Notes Chemistry Chapter Wise Notes
    - Highlighting and Note-Taking Class 12th Notes Chemistry Chapter Wise Notes
    - Interactive Elements Class 12th Notes Chemistry Chapter Wise Notes
  8. Staying Engaged with Class 12th Notes Chemistry Chapter Wise Notes
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Class 12th Notes Chemistry Chapter Wise Notes
  9. Balancing eBooks and Physical Books Class 12th Notes Chemistry Chapter Wise Notes
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Class 12th Notes Chemistry Chapter Wise Notes
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Class 12th Notes Chemistry Chapter Wise Notes
    - Setting Reading Goals Class 12th Notes Chemistry Chapter Wise Notes
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Class 12th Notes Chemistry Chapter Wise Notes
    - Fact-Checking eBook Content of Class 12th Notes Chemistry Chapter Wise Notes
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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