

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? **copper (II) chloride**

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction? Percent yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

Limiting Reactant And Percent Yield Answers

LL Leslie



Limiting Reactant And Percent Yield Answers:

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Prentice Hall Chemistry, 2000

Preparatory Chemistry Howard Stephen Stoker, 1990

Introduction to Chemical Principles Howard Stephen Stoker, 1990 Introduction to Chemical Principles is a text for students who have had little to no previous instruction in chemistry or who had such instruction long enough ago that a thorough review is needed.

preface

College Chemistry Harold Goldwhite, John R. Spielman, 1984 An outline of the basic

concepts of chemistry includes discussions of scientific notation atomic structure chemical bonding and the periodic table

Improving Student Comprehension in Chemistry Laboratories Tracy Lynn Haroff, 2006 **Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb** Wade A. Freeman, 1986 **Student Study Guide to accompany Chemistry** Martin Silberberg, Libby Weberg, 2005-01-06 **Chemistry 2007** David Wilson, 2006-12 Kaplan's guide includes 2 full length practice tests Diagnostic test to target areas for score improvement Detailed answer explanations Hundreds of practice questions from calculations of chemical equations to organic chemistry Explanations of important terms formulas and concepts Powerful strategies to help you score higher *Basic Chemical Principles* Donada Peters, Edward I. Peters, 1988 **Chemistry** John Christian Bilar, 1989 Tb Chemistry Molecular Sci Moore, 2001-12 **Selected Solutions for Chemistry, Concepts and Models by Robinson, Odom, and Holtzclaw** John H. Meiser, 1992 *Chemistry* Zumdahl, Steven S. Zumdahl, 2003-07 Contains chapter discussions key term definitions and practice chapter tests **5 Steps to a 5: 500 AP Chemistry Questions to Know by Test Day, Third Edition** Mina Lebitz, 2019-02-21 500 AP style questions with detailed answer explanations to prepare you for what you'll see on test day 5 Steps to a 5 500 AP Chemistry Questions to Know by Test Day gives you 500 practice questions that cover the most essential course material and help you work toward a 5 on the test The questions parallel the format and degree of difficulty that you'll find on the actual AP exams and are accompanied by answers with comprehensive explanations The questions in this book were written by expert AP teachers who know the exam inside and out so they closely reflect what you'll see when you sit for the AP Chemistry test This valuable study guide features 500 AP style questions and answers Detailed review explanations for right and wrong answers Close simulations of the real AP exam Updated material that reflects the latest AP exam **McGraw-Hill Education 500 AP Chemistry Questions to Know by Test Day, 2nd edition** Mina Lebitz, 2015-10-30 Get ready for your AP Chemistry exam with 500 AP Chemistry questions updated for all the latest exam changes We want you to succeed on your AP exams That's why we've selected these 500 questions to help you study more effectively use your preparation time wisely and get your best score These AP style questions and answers are similar to the ones you'll find on the exam so you will know what to expect on your test day Each question includes detailed explanation with right and wrong answers to enhance your full understanding of the concepts Whether you are just beginning your test preparation or doing a last minute review 5 Steps to a 5 500 AP Chemistry Questions 2ed will help you achieve the score you desire 500 AP style questions and answers Complete answer explanations for every question What you really need to know to achieve a high score *Introductory Chemistry* Nivaldo J. Tro, 2006

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