

CHAPTER

02

SOLUTION

1. **Solution** : A homogeneous mixture of two or more substances is known as solution.
2. **Solute** : The substance present in smaller amount in a solution is called solute.
3. **Solvent** : The substance present in larger amount in a solution is called solvent.
4. **Binary solution** : These solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
5. **Ternary solutions** : Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
6. **Aqueous solution** : When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
7. **Non-aqueous solution** : When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
8. Weight Fraction = $\frac{\text{weight of solute}}{\text{weight of solution}}$
9. Weight percentage

$$= \frac{\text{weight of solute}}{\text{weight of solution}} \times 100$$

Weight of solute present in 100 grams of solution is called weight percentage

Volume of solute present in 100 ml of the solution is called volume percentage
10. Volume percentage

$$= \frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$$
11. Volume Fraction = $\frac{\text{Volume of solute}}{\text{Volume of solution}}$
12. Molarity (M)

$$= \frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$$

(or) $M = \frac{W}{GMW} \times \frac{1}{V(lit)}$
13. Molarity = $\frac{\text{No of moles of solute}}{\text{Volume of solution in ml}} \times 1000$
14. Molarity = $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{V \text{ in ml}}$
 (or) $M = \frac{W}{GMW} \times \frac{1000}{V(ml)}$
15. Weight of solute = $\frac{M \times GMW \times V \text{ in ml}}{1000}$
16. Molarity for dilute solutions : $M_1V_1 = M_2V_2$
17. For neutralisation reactions : $\frac{M_1V_1}{n_1} = \frac{M_2V_2}{n_2}$
18. For mixed solution molarity

$$= \frac{M_1V_1 + M_2V_2 + \dots}{V_1 + V_2 + \dots}$$

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Table of Contents Class Xii Chemistry Ch 2 Solutions

1. Understanding the eBook Class Xii Chemistry Ch 2 Solutions
 - The Rise of Digital Reading Class Xii Chemistry Ch 2 Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Class Xii Chemistry Ch 2 Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Class Xii Chemistry Ch 2 Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Class Xii Chemistry Ch 2 Solutions
 - Personalized Recommendations
 - Class Xii Chemistry Ch 2 Solutions User Reviews and Ratings
 - Class Xii Chemistry Ch 2 Solutions and Bestseller Lists
5. Accessing Class Xii Chemistry Ch 2 Solutions Free and Paid eBooks

- Class Xii Chemistry Ch 2 Solutions Public Domain eBooks
 - Class Xii Chemistry Ch 2 Solutions eBook Subscription Services
 - Class Xii Chemistry Ch 2 Solutions Budget-Friendly Options
6. Navigating Class Xii Chemistry Ch 2 Solutions eBook Formats
- ePub, PDF, MOBI, and More
 - Class Xii Chemistry Ch 2 Solutions Compatibility with Devices
 - Class Xii Chemistry Ch 2 Solutions Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Class Xii Chemistry Ch 2 Solutions
 - Highlighting and Note-Taking Class Xii Chemistry Ch 2 Solutions
 - Interactive Elements Class Xii Chemistry Ch 2 Solutions
8. Staying Engaged with Class Xii Chemistry Ch 2 Solutions
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Class Xii Chemistry Ch 2 Solutions
9. Balancing eBooks and Physical Books Class Xii Chemistry Ch 2 Solutions
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Class Xii Chemistry Ch 2 Solutions
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Class Xii Chemistry Ch 2 Solutions
- Setting Reading Goals Class Xii Chemistry Ch 2 Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Class Xii Chemistry Ch 2 Solutions
- Fact-Checking eBook Content of Class Xii Chemistry Ch 2 Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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