

**Assignment #2 – Compound Names and Formulas**  
**Multi-valent ions only**

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

- |                              |                                    |                            |                                    |
|------------------------------|------------------------------------|----------------------------|------------------------------------|
| 1. copper (II) oxide         | <u>CuO</u>                         | 11. manganese (III) oxide  | <u>Mn<sub>2</sub>O<sub>3</sub></u> |
| 2. mercury (I) oxide         | <u>Hg<sub>2</sub>O</u>             | 12. vanadium (II) bromide  | <u>VBr<sub>2</sub></u>             |
| 3. gold (III) chloride       | <u>AuCl<sub>3</sub></u>            | 13. niobium(V) oxide       | <u>Nb<sub>2</sub>O<sub>5</sub></u> |
| 4. thallium (III) bromide    | <u>TlBr<sub>3</sub></u>            | 14. titanium (III) oxide   | <u>Ti<sub>2</sub>O<sub>3</sub></u> |
| 5. bismuth (V) oxide         | <u>Bi<sub>2</sub>O<sub>5</sub></u> | 15. titanium (III) nitride | <u>TiN</u>                         |
| 6. terbium (III) oxide       | <u>Tb<sub>2</sub>O<sub>3</sub></u> | 16. iron (II) oxide        | <u>FeO</u>                         |
| 7. uranium (VI) oxide        | <u>UO<sub>3</sub></u>              | 17. cobalt (II) phosphide  | <u>Co<sub>3</sub>P<sub>2</sub></u> |
| 8. protactinium (IV) bromide | <u>PaBr<sub>4</sub></u>            | 18. tin (II) oxide         | <u>SnO</u>                         |
| 9. cerium (III) oxide        | <u>Ce<sub>2</sub>O<sub>3</sub></u> | 19. thulium (II) bromide   | <u>TmBr<sub>2</sub></u>            |
| 10. arsenic (V) sulphide     | <u>As<sub>2</sub>S<sub>5</sub></u> | 20. copper (II) bromide    | <u>CuBr<sub>2</sub></u>            |

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

- |                                   |                             |                                    |                                |
|-----------------------------------|-----------------------------|------------------------------------|--------------------------------|
| 1. SnCl <sub>4</sub>              | <u>Tin (IV) Chloride</u>    | 9. PdF <sub>4</sub>                | <u>Palladium (IV) Fluoride</u> |
| 2. BiBr <sub>5</sub>              | <u>Bismuth (V) Bromide</u>  | 10. Os <sub>2</sub> O <sub>3</sub> | <u>Osmium (III) Oxide</u>      |
| 3. PoO <sub>2</sub>               | <u>Polonium (IV) Oxide</u>  | 11. MoBr <sub>2</sub>              | <u>Molybdenum (II) Bromide</u> |
| 4. PbI <sub>2</sub>               | <u>Lead (II) Iodide</u>     | 12. VCl <sub>5</sub>               | <u>Vanadium (V) Chloride</u>   |
| 5. HgO                            | <u>Mercury (II) Oxide</u>   | 13. Mn <sub>2</sub> O <sub>3</sub> | <u>Manganese (III) Oxide</u>   |
| 6. HgCl                           | <u>Mercury (I) Chloride</u> | 14. CoO                            | <u>Cobalt (II) Oxide</u>       |
| 7. Au <sub>2</sub> O <sub>3</sub> | <u>Gold (III) Oxide</u>     | 15. Np <sub>2</sub> O <sub>3</sub> | <u>Neptunium (III) Oxide</u>   |
| 8. FeCl <sub>2</sub>              | <u>Iron (II) Chloride</u>   | 16. V <sub>2</sub> O <sub>5</sub>  | <u>Vanadium (V) Oxide</u>      |

# Nomenclature And Formula Writing Worksheet Answers

**Jiyuan Zhang**



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## **Table of Contents Nomenclature And Formula Writing Worksheet Answers**

1. Understanding the eBook Nomenclature And Formula Writing Worksheet Answers
  - The Rise of Digital Reading Nomenclature And Formula Writing Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Nomenclature And Formula Writing Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Nomenclature And Formula Writing Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Nomenclature And Formula Writing Worksheet Answers
  - Personalized Recommendations
  - Nomenclature And Formula Writing Worksheet Answers User Reviews and Ratings

- Nomenclature And Formula Writing Worksheet Answers and Bestseller Lists
- 5. Accessing Nomenclature And Formula Writing Worksheet Answers Free and Paid eBooks
  - Nomenclature And Formula Writing Worksheet Answers Public Domain eBooks
  - Nomenclature And Formula Writing Worksheet Answers eBook Subscription Services
  - Nomenclature And Formula Writing Worksheet Answers Budget-Friendly Options
- 6. Navigating Nomenclature And Formula Writing Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Nomenclature And Formula Writing Worksheet Answers Compatibility with Devices
  - Nomenclature And Formula Writing Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Nomenclature And Formula Writing Worksheet Answers
  - Highlighting and Note-Taking Nomenclature And Formula Writing Worksheet Answers
  - Interactive Elements Nomenclature And Formula Writing Worksheet Answers
- 8. Staying Engaged with Nomenclature And Formula Writing Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Nomenclature And Formula Writing Worksheet Answers
- 9. Balancing eBooks and Physical Books Nomenclature And Formula Writing Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Nomenclature And Formula Writing Worksheet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Nomenclature And Formula Writing Worksheet Answers
  - Setting Reading Goals Nomenclature And Formula Writing Worksheet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nomenclature And Formula Writing Worksheet Answers
  - Fact-Checking eBook Content of Nomenclature And Formula Writing Worksheet Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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