

## CHAPTER 8 REVIEW

**Chemical Equations and Reactions****SHORT ANSWER** Answer the following questions in the space provided.

2. \_\_\_\_\_ According to the law of conservation of mass, the total mass of the reacting substances is
- always more than the total mass of the products.
  - always less than the total mass of the products.
  - sometimes more and sometimes less than the total mass of the products.
  - always equal to the total mass of the products.**

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 2. In the reaction described by the word equation  

$$\text{sodium} + \text{water} \rightarrow \text{sodium hydroxide} + \text{hydrogen gas}$$
  - sodium hydroxide is a product.
  - sodium is a product.
  - water is a reactant.
  - Both (a) and (c).**
- \_\_\_\_\_ 3. Which of the following indicates that a chemical equation is balanced?  
  - The numbers of atoms of each element are the same on both sides of the equation.**
  - All of the coefficients are the same.
  - The numbers of molecules on each side of the equation are equal.
  - The sums of the coefficients on each side of the equation are equal.
- \_\_\_\_\_ 4. From a complete and correctly written chemical equation, you can obtain the  
  - chemical formulas of the reactants and products.
  - relative amounts of the reactants and products.
  - physical states of the reactants and products.
  - All of the above.**
- \_\_\_\_\_ 5. To balance a chemical equation, you must  
  - adjust the subscripts.
  - adjust the number of products.
  - adjust the number of reactants.
  - adjust the coefficients.**
- \_\_\_\_\_ 6. The formation of a solid product in a chemical reaction is represented by the symbol  
  - (aq).
  - (g).
  - (l).
  - (s).**
- \_\_\_\_\_ 8. In the expression  $3\text{CO}_2$ , the numbers 3 and 2 are, respectively,  
  - a subscript and a coefficient.
  - a coefficient and a subscript.**
  - two subscripts.
  - two coefficients.
- \_\_\_\_\_ 9. What is the total number of molecules that is represented by the following chemical equation? (Add up the coefficients.)  

$$\text{NH}_4\text{NO}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$$
  - 3
  - 4**
  - 8
  - 16
- \_\_\_\_\_ 10. Which equation below violates the law of conservation of mass?  
  - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
  - $\text{KCl} + \text{Br} \rightarrow \text{KBr} + \text{Cl}_2$**
  - $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
  - $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$
- \_\_\_\_\_ 11. Which of the following equations represents the balanced equation for the reaction of iron and oxygen?  
  - $2\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
  - $\text{Fe} + 3\text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
  - $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$**
  - $3\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
- \_\_\_\_\_ 12. Which coefficients correctly balance the formula equation below?  

$$\text{KClO}_3(\text{s}) \rightarrow \text{KCl}(\text{s}) + \text{O}_2(\text{g})$$
  - 1, 1, 1
  - 1, 1, 3
  - 2, 2, 3**
  - 2, 1, 1
- \_\_\_\_\_ 13. Which equation is *not* balanced?  
  - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
  - $4\text{H}_2 + 2\text{O}_2 \rightarrow 4\text{H}_2\text{O}$
  - $\text{H}_2 + \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$
  - $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$**
- \_\_\_\_\_ 14. In which type of reaction do two or more compounds react to form one product?  
  - synthesis (AKA combination)**
  - decomposition
  - combustion
  - single-replacement
- \_\_\_\_\_ 15. Which of the following is true regarding single-replacement reactions?  
  - Only one product is formed.
  - There is only one reagent.
  - It requires oxygen.
  - One element replaces another in a compound.**

# Chapter 8 Review Answers Chemistry

**JE Gale**



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**hydraulische presse konstruktion einsatzgebiete und größen** - Sep 04 2022

web die hydraulische presse ist ein bearbeitungswerkzeug welches eingesetzt wird wenn sehr große kräfte auf ein werkstück eingebracht werden sollen sie kann hervorragende umformarbeiten leisten ihr nachteil ist dass sie sehr langsam arbeitet die hydraulische presse wird daher vorwiegend zur bearbeitung von einzelstücken verwendet

*hydraulische pressen pressanlagen von der schubert group* - Feb 26 2022

web hydraulische pressen auch bekannt als hydraulik pressen verfügen über hohe presskräfte bei relativ geringer pressgeschwindigkeit also einer geringen geschwindigkeit des presskolbens Über den gesamten pressweg gewährleisten hydraulische pressen eine gleichmäßige kraftverteilung was ein hohes maß an präzision bei sämtlichen press

hydraulische pressen sonstige umformmaschinen hydac - Dec 07 2022

web um die energie und ressourceneffizienz von hydraulischen pressen zu verbessern setzen wir auf bauraumoptimierte hydrauliktanks sowie bedarfsgerechter energiebereitstellung profitieren sie von unserem umfassenden produktsortiment von der komponente bis zum system wir begleiten sie auf dem weg in eine nachhaltigere zukunft

*die hydraulische presse was ist das und wie funktioniert sie* - Jul 14 2023

web feb 13 2023 was ist die hydraulische presse eine hydraulische presse ist die einfachste hydraulische maschine die entwickelt wurde um signifikante druckkräfte auf der grundlage des pascalschen gesetzes zu erzielen früher brahm presse genannt da sie 1795 von joseph bramah erfunden und patentiert wurde