

## Unit 11: Chapter 7 - Chemical Reactions Review Answer Key

**Define:**

1. **Reactant:** *A substance that undergoes change in a chemical reaction*
2. **Product:** *New substances formed as a result of a chemical reaction.*
3. **Catalyst:** *A substance that affects the rate of a chemical reaction without being used up in the reaction.*
4. **Coefficient:** *Numbers that appear before a formula in a chemical equation to show the relative proportions of each reactant and product*
5. **Subscript:** *A number written lower in a chemical formula indicating the amount of atoms present of an element (this number is not to be changed once written in chemical formula)*
6. **Law of Conservation of Mass:** *Matter cannot be created nor destroyed; only changed (atoms can be rearranged)*
7. **Synthesis reaction (include an example):** *A chemical reaction in which two or more substances react to form a single substance.*       $(A + B \rightarrow AB)$
8. **Decomposition reaction (include an example):** *A chemical reaction in which a compound breaks down into two or more simpler substances.*       $(AB \rightarrow A + B)$
9. **Single replacement (include an example):** *A chemical reaction in which one element takes the place of another element in a compound.*       $(AB + C \rightarrow AC + B)$
10. **Double replacement (include an example):** *A chemical reaction in which two compounds exchange positive ions and form two new compounds.*       $(AB + CD \rightarrow AD + CB)$
11. **Combustion (include an example):** *A chemical reaction in which a substance reacts rapidly with oxygen, often producing heat and light.*       $(C_2H_2 + O_2 \rightarrow CO_2 + H_2O)$
12. **Endothermic reaction:** *A chemical reaction that **absorbs** energy from its surroundings.*
13. **Exothermic reaction:** *A chemical reaction that **releases** energy to its surroundings.*
14. **Chemical equation (include an example):** *A representation of a chemical reaction in which the reactants and products are expressed as formulas.*
15. A **balanced equation** is an equation in which the number of atoms on the reactants' side **equals** the number of atoms on the products' side of the chemical equation.
16. **Equilibrium:** *A state in which the forward and reverse paths of a physical or chemical change take place at the same rate.*

# Chapter 7 Chemical Reactions Answer Key

**SA Dillow**



## **Chapter 7 Chemical Reactions Answer Key:**

Chemistry (Teacher Guide) Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter, helps each student stay on schedule and be organized, and is their source of accountability along the way. With that in mind, this guide provides additional help through the laboratory exercises as well as lessons, quizzes, and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study. Features: Each suggested weekly schedule has five easy-to-manage lessons that combine reading and worksheets. Worksheets, quizzes, and tests are perforated and three-hole punched; materials are easy to tear out, hand out, grade, and store. Adjust the schedule and materials needed to best work within your educational program. Space is given for assignments and dates. There is flexibility in scheduling. Adapt the days to your school schedule. Workflow: Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade. About the Author: DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and Au Sable Institute of Environmental Studies.

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enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place™ Special Edition and on The Chemistry of Life CD ROM For college instructors and students *Books in Print Supplement*, 2002 General and Inorganic Chemistry Irena Kostova, 2025-04-21

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## **Table of Contents Chapter 7 Chemical Reactions Answer Key**

1. Understanding the eBook Chapter 7 Chemical Reactions Answer Key
  - The Rise of Digital Reading Chapter 7 Chemical Reactions Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 7 Chemical Reactions Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 7 Chemical Reactions Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 7 Chemical Reactions Answer Key
  - Personalized Recommendations
  - Chapter 7 Chemical Reactions Answer Key User Reviews and Ratings
  - Chapter 7 Chemical Reactions Answer Key and Bestseller Lists
5. Accessing Chapter 7 Chemical Reactions Answer Key Free and Paid eBooks

- Chapter 7 Chemical Reactions Answer Key Public Domain eBooks
  - Chapter 7 Chemical Reactions Answer Key eBook Subscription Services
  - Chapter 7 Chemical Reactions Answer Key Budget-Friendly Options
6. Navigating Chapter 7 Chemical Reactions Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Chapter 7 Chemical Reactions Answer Key Compatibility with Devices
    - Chapter 7 Chemical Reactions Answer Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chapter 7 Chemical Reactions Answer Key
    - Highlighting and Note-Taking Chapter 7 Chemical Reactions Answer Key
    - Interactive Elements Chapter 7 Chemical Reactions Answer Key
  8. Staying Engaged with Chapter 7 Chemical Reactions Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chapter 7 Chemical Reactions Answer Key
  9. Balancing eBooks and Physical Books Chapter 7 Chemical Reactions Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chapter 7 Chemical Reactions Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chapter 7 Chemical Reactions Answer Key
    - Setting Reading Goals Chapter 7 Chemical Reactions Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chapter 7 Chemical Reactions Answer Key
    - Fact-Checking eBook Content of Chapter 7 Chemical Reactions Answer Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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