

## CHAPTER 3 Matter—Properties and Changes

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### 3.1 Properties of Matter

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All of the material—the “stuff”—around us is matter. A **substance** is matter that has a uniform and unchanging composition. For example, water is a pure substance. No matter where it is found, a sample of water will have the same composition as any other sample of water.

A **physical property** of a substance is a characteristic that can be observed and measured without changing the composition of the substance. Words such as *hard*, *soft*, *shiny*, *dull*, *brittle*, *flexible*, *heavy* (in density), and *light* (in density) are used to describe physical properties.

A **chemical property** describes the ability of a substance to combine with or change into one or more other substances. For example, the ability of iron to form rust when combined with air is a chemical property of iron. The inability of a substance to combine with another substance is also a chemical property. For example, the inability to combine with most other substances is a chemical property of gold.

#### Practice Problems

1. Identify each of the following as an example of a physical property or a chemical property.
  - a. Silver tarnishes when it comes in contact with hydrogen sulfide in the air.
  - b. A sheet of copper can be pounded into a bowl.
  - c. Barium melts at  $725^{\circ}\text{C}$ .
  - d. Helium does not react with any other element.
  - e. A bar of lead is more easily bent than is a bar of aluminum of the same size.
  - f. Potassium metal is kept submerged in oil to prevent contact with oxygen or water.

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