

5210 B. 5-Day BOD Test

1. General Discussion

The BOD test is an indirect measurement of organic matter; it measures the change in DO concentration caused by microorganisms as they degrade organic matter in a sample held in a stoppered bottle incubated for 5 d in the dark at 20°C. Analysts measure DO before and after incubation, and compute BOD using the difference between DO measurements. Because initial DO is determined shortly after dilution, all oxygen uptake occurring after this measurement is included in the BOD measurement.

For sampling and storage procedures, see 5210B.4a.

2. Apparatus

a. Incubation bottles: Use 60-mL, glass bottles or larger (300-mL bottles with a flared mouth and ground-glass stopper are preferred). Clean bottles with a detergent, rinse thoroughly, and drain before use. Alternatively, use disposable plastic BOD bottles that are capable of meeting all method quality-control (QC) checks.

b. Air incubator or water bath, thermostatically controlled at 20 ± 1°C. Exclude all light to prevent the possibility of photosynthetic production of DO.

c. Oxygen-sensitive membrane electrode, polarographic or galvanic, or *oxygen-sensitive optical probe* with appropriate meter.

3. Reagents

Discard reagents if there is any sign of precipitation or biological growth in the stock bottles. Commercial equivalents of these reagents are acceptable, and different stock concentrations may be used if doses are adjusted proportionally. Use reagent grade or better for all chemicals and use distilled or equivalent reagent-grade water (see Section 1060) to make all solutions.

a. Phosphate buffer solution: Dissolve 8.5 g monopotassium phosphate (KH_2PO_4), 21.75 g dipotassium phosphate (K_2HPO_4), 33.4 g disodium phosphate ($\text{Na}_2\text{HPO}_4 \cdot 7\text{H}_2\text{O}$), and 1.7 g ammonium chloride (NH_4Cl) in about 500 mL reagent-grade water and dilute to 1 L. The pH should be 7.2 without further adjustment. Alternatively, dissolve 42.5 g KH_2PO_4 and 1.7 g NH_4Cl in about 700 mL reagent-grade water. Adjust pH to 7.2 with 30% sodium hydroxide (NaOH) and dilute to 1 L.

b. Magnesium sulfate (MgSO_4) solution: Dissolve 22.5 g $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$ in reagent-grade water and dilute to 1 L.

c. Calcium chloride (CaCl_2) solution: Dissolve 27.5 g CaCl_2 in reagent-grade water and dilute to 1 L.

d. Ferric chloride (FeCl_3) solution: Dissolve 0.25 g $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ in reagent-grade water and dilute to 1 L.

e. Acid and alkali solutions, 1N, to neutralize caustic or acidic waste samples.

1) *Acid*—Slowly and while stirring, add 28 mL conc sulfuric acid (H_2SO_4) to reagent-grade water. Dilute to 1 L.

2) *Alkali*—Dissolve 40 g NaOH in distilled water. Dilute to 1 L.

f. Sodium sulfite (Na_2SO_3) solution: Dissolve 1.575 g Na_2SO_3 in 1000 mL reagent-grade water. This solution is unstable; prepare daily.

g. Nitrofurantoin inhibitor:

1) 2-chloro-6-(trichloromethyl) pyridine (TCMP)—Use pure TCMP or commercial preparations.*

2) Allylthiourea (ATU) solution—Dissolve 2.0 g allylthiourea ($\text{C}_4\text{H}_9\text{N}_2\text{S}$) in about 500 mL reagent-grade water and dilute to 1 L. Store at 4°C. The solution is stable for 2 weeks when stored at 6°C without freezing.

h. Glucose-glutamic acid (GGA) solution: Dry reagent-grade glucose and reagent-grade glutamic acid at 103°C for 1 h. Add 150 mg glucose and 150 mg glutamic acid to reagent-grade water and dilute to 1 L. Prepare fresh immediately before use unless solution is maintained in a sterile container. Store all GGA mixtures at 6°C without freezing unless manufacturer recommendations state otherwise. Commercial preparations may be used but concentrations may vary. Discard solutions if evi-

dence of contamination is indicated (e.g., growth occurs in the stock bottle or GGA test results are consistently low).

i. Ammonium chloride solution: Dissolve 1.15 g NH_4Cl in about 500 mL reagent-grade water, adjust pH to 7.2 with NaOH solution, and dilute to 1 L. Solution contains 0.3 mg N/mL.

j. Source water for preparing BOD dilution water: Use demineralized, distilled, or equivalent reagent-grade water, tap, or natural water to make sample dilutions (see 5210B.4c).

4. Preparatory Procedures

a. Sampling and storage: Samples for BOD analysis may degrade significantly during storage between collection and analysis, resulting in low BOD values.

1) *Grab samples*—If analysis begins within 2 h of collection, cold storage is unnecessary. Otherwise, keep sample at 6°C between collection and analysis. Ideally, begin analysis within 6 h of sample collection; if impossible due to distance between sampling site and laboratory, then begin analysis within 24 h of collection. The recommended holding time is 24 h; however, the U.S. Environmental Protection Agency (EPA) allows for a 48-h holding time.

2) *Composite samples*—Limit compositing period to 24 h, and keep samples at 6°C during process. Store for the same time and temperature as grab samples, although in this case, holding time begins when the compositing period ends.

b. Sample preparation and pretreatment:

1) *All samples*—Check pH; if it is not between 6.0 and 8.0, adjust sample temperature to 20 ± 3°C, then adjust pH to between 6.5 and 7.5 using an H_2SO_4 or NaOH solution strong enough not

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Third Century of Biochemical Oxygen Demand Rodger Baird, Roy Keith Smith, 2002 The most comprehensive summary and literature review of Biochemical Oxygen Demand BOD on the market BOD is one of the fundamental concepts in wastewater treatment Throughout the 1800s and the 1900s BOD was exhaustively studied and refined both as a concept and as an analytical procedure Review all previous BOD work including why technicians scientists plant operators regulators and engineers have complained about the BOD test for many years This book is intended to serve three purposes first and foremost is to describe BOD as a test procedure and biological phenomenon secondly to describe the place of BOD within the complex of testing that is used to evaluate treatment processes lastly is to present the development of BOD and preserve all peer reviewed literature citations that mark the road to the current test Written by Rodger B Baird and Roy Keith Smith Mr Baird is Manager of Laboratories for the Los Angeles County Sanitation Districts where he has worked in laboratory supervision and management for more than 30 years Dr Smith has written over 60 published articles and is the author of eight books one in the 4th edition on environmental analysis Chapters detailing sediment oxygen demand chemical oxygen demand and total organic carbon testing and their relationship to BOD testing as well as extensive coverage of the interferences encountered during oxygen demand testing makes this a must have reference *Appita Journal*, 1998

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Conference on Water for Texas ,1995 *Introduction to Environmental Engineering* Mackenzie Davis,David A. Cornwell,2006-10-03 Introduction to Environmental Engineering 4 e contains the essential science and engineering principles needed for introductory courses and used as the basis for more advanced courses in environmental engineering Updated with latest EPA regulations Davis and Cornwell apply the concepts of sustainability and materials and energy balance as a means of understanding and solving environmental engineering issues With 650 end of chapter problems as well as provocative discussion questions and a helpful list of review items found at the end of each chapter the text is both a comprehensible and comprehensive tool for any environmental engineering course Standards and Laws are the most current and up to date for an environmental engineering text **Transactions of the Kentucky Academy of Science** Kentucky Academy of Science,1996 List of members in v 1 Rotenone Use in Fisheries Management Brian J. Finlayson,2000 □□□ □□□□ ,1998 *Standard Methods for the Examination of Water and Wastewater* APHA-AWWA-WPCF.,1981 **Standard Methods for the Examination of Water and Wastewater** Apha,1985 *Standard Methods for the Examination of Water and Wastewater 14ed* APHA.,1976

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Table of Contents Apha Bod 5210b

1. Understanding the eBook Apha Bod 5210b
 - The Rise of Digital Reading Apha Bod 5210b
 - Advantages of eBooks Over Traditional Books
2. Identifying Apha Bod 5210b
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Apha Bod 5210b
 - User-Friendly Interface
4. Exploring eBook Recommendations from Apha Bod 5210b
 - Personalized Recommendations
 - Apha Bod 5210b User Reviews and Ratings

- Apha Bod 5210b and Bestseller Lists
- 5. Accessing Apha Bod 5210b Free and Paid eBooks
 - Apha Bod 5210b Public Domain eBooks
 - Apha Bod 5210b eBook Subscription Services
 - Apha Bod 5210b Budget-Friendly Options
- 6. Navigating Apha Bod 5210b eBook Formats
 - ePub, PDF, MOBI, and More
 - Apha Bod 5210b Compatibility with Devices
 - Apha Bod 5210b Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Apha Bod 5210b
 - Highlighting and Note-Taking Apha Bod 5210b
 - Interactive Elements Apha Bod 5210b
- 8. Staying Engaged with Apha Bod 5210b
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Apha Bod 5210b
- 9. Balancing eBooks and Physical Books Apha Bod 5210b
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Apha Bod 5210b
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Apha Bod 5210b
 - Setting Reading Goals Apha Bod 5210b
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Apha Bod 5210b
 - Fact-Checking eBook Content of Apha Bod 5210b
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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