

Acids and Bases Worksheet 1

1. Finish the following reactions when acids are added to water.



2. What are the conjugate bases of these acids?

| original acid | conjugate base |
|-------------------------|----------------|
| HNO_3 | |
| H_2O | |
| H_3O^+ | |
| H_2SO_4 | |
| HBr | |
| HCO_3^- | |

3. What are the conjugate acids of these bases?

| original base | conjugate acid |
|----------------------|----------------|
| OH^- | |
| H_2O | |
| HCO_3^- | |
| SO_4^{2-} | |
| ClO_4^- | |

4. Which of the following represent conjugate acid-base pairs?



5. Calculate the $[\text{H}^+]$ in a solution in which $[\text{OH}^-] = 2.0 \times 10^{-2} \text{ M}$. Is this solution acidic, neutral, or basic?

6. Calculate the $[\text{OH}^-]$ in a solution in which $[\text{H}^+] = 3.99 \times 10^{-5} \text{ M}$. Is this solution acidic, neutral, or basic?

Acids And Bases Study Guide Answer Key

L Darling-Hammond



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