

9/2/16

S. Sai Krishna

Date _____
Page _____
ClassmateThermochemistry

$$Q = \Delta E - W$$

$$Q = \Delta E - (-P\Delta V)$$

$$Q = \Delta E + P\Delta V$$

- At constant Volume.

$$Q = \Delta E$$

- At const. Pressure.

$$Q_p = \Delta E + P\Delta V$$

$$H = \Delta E + P\Delta V$$

$$H = \Delta E + \Delta nRT$$

$\Delta n =$ no. of moles of Prod
- no. of moles of react.

→ For ΔH .

- If ΔH is -ve. reaction is exothermic.
- If ΔH is +ve. reaction is endothermic.

Q. The heat of combustion of ethane at 17°C and const volume is -322.1 kJ . Calculate the heat of combustion at constant pressure.



Engineering Chemistry 2 Notes

LL Leslie



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40 g of H_2O molar mass of H_2O is 18 gm mole and molar mass of CH_3OH is 32 gm mole solution moles of CH_3OH 4.5 32 0 14 mole moles of H_2O 40 18 2 2 moles therefore according to the equation

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