

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative	noble gas	octet	electron configuration
positive	electrons	ions	polyatomic
crisscross	anion	cation	empirical formula

- An ionic bond is an attraction between oppositely charged ions.
- Anions have a negative charge.
- An atom becomes an ion by losing or gaining electrons.
- The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
- When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
- The atoms that make up polyatomic ions are bonded together by covalent bonds.
- The crisscross method can be used to write the formula for an ionic compound.
- The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- | | |
|------------------------|--|
| <u>losing</u> | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons. |
| <u>true</u> | 10. A cation has a <u>positive</u> charge. |
| <u>true</u> | 11. A binary ionic compound contains only <u>one</u> kind of cation and one kind of anion. |
| <u>polyatomic</u> | 12. <u>Monatomic</u> ions consist of more than one atom. |
| <u>seven</u> | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons. |
| <u>true (released)</u> | 14. A <u>great deal of energy is produced</u> when an ionic compound is formed. |

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M Lipman



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Table of Contents Grade 7 Chapter Test Answer Keys Prentice Hall

1. Understanding the eBook Grade 7 Chapter Test Answer Keys Prentice Hall
 - The Rise of Digital Reading Grade 7 Chapter Test Answer Keys Prentice Hall
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 7 Chapter Test Answer Keys Prentice Hall
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 7 Chapter Test Answer Keys Prentice Hall
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 7 Chapter Test Answer Keys Prentice Hall
 - Personalized Recommendations
 - Grade 7 Chapter Test Answer Keys Prentice Hall User Reviews and Ratings
 - Grade 7 Chapter Test Answer Keys Prentice Hall and Bestseller Lists
5. Accessing Grade 7 Chapter Test Answer Keys Prentice Hall Free and Paid eBooks

- Grade 7 Chapter Test Answer Keys Prentice Hall Public Domain eBooks
 - Grade 7 Chapter Test Answer Keys Prentice Hall eBook Subscription Services
 - Grade 7 Chapter Test Answer Keys Prentice Hall Budget-Friendly Options
6. Navigating Grade 7 Chapter Test Answer Keys Prentice Hall eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade 7 Chapter Test Answer Keys Prentice Hall Compatibility with Devices
 - Grade 7 Chapter Test Answer Keys Prentice Hall Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 7 Chapter Test Answer Keys Prentice Hall
 - Highlighting and Note-Taking Grade 7 Chapter Test Answer Keys Prentice Hall
 - Interactive Elements Grade 7 Chapter Test Answer Keys Prentice Hall
 8. Staying Engaged with Grade 7 Chapter Test Answer Keys Prentice Hall
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 7 Chapter Test Answer Keys Prentice Hall
 9. Balancing eBooks and Physical Books Grade 7 Chapter Test Answer Keys Prentice Hall
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 7 Chapter Test Answer Keys Prentice Hall
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Grade 7 Chapter Test Answer Keys Prentice Hall
 - Setting Reading Goals Grade 7 Chapter Test Answer Keys Prentice Hall
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Grade 7 Chapter Test Answer Keys Prentice Hall
 - Fact-Checking eBook Content of Grade 7 Chapter Test Answer Keys Prentice Hall
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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