

Chapter 17: Chemical Equilibrium

Worksheet 17.1: K_c and K_p

1. Consider the following reaction:

- 1. $2\text{NO}_2(\text{g}) \rightleftharpoons \text{N}_2\text{O}_4(\text{g})$ $K_c = 5.0$ at 300 K
- 2. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ $K_c = 0.50$ at 400 K
- 3. $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$ $K_c = 50.0$ at 400 K
- 4. $\text{CO}_2(\text{g}) + \text{H}_2(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g})$ $K_c = 1.0$ at 1000 K
- 5. $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$ $K_c = 4.0$ at 1000 K
- 6. $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2\text{HCl}(\text{g})$ $K_c = 4.0$ at 1000 K
- 7. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ $K_c = 0.50$ at 400 K
- 8. $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$ $K_c = 50.0$ at 400 K
- 9. $\text{CO}_2(\text{g}) + \text{H}_2(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g})$ $K_c = 1.0$ at 1000 K
- 10. $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$ $K_c = 4.0$ at 1000 K

Answers

- 1. $K_c = 0.20$
- 2. $K_c = 4.0$
- 3. $K_c = 0.020$
- 4. $K_c = 1.0$
- 5. $K_c = 0.25$
- 6. $K_c = 0.25$
- 7. $K_c = 0.50$
- 8. $K_c = 50.0$
- 9. $K_c = 1.0$
- 10. $K_c = 0.25$

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