

# Atomic structure and the periodic table

## Atomic structure



protons  
neutrons  
electrons

atoms have no overall charge

particle	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1

## elements

elements - substances made up of atoms with the same atomic number

isotopes of an element - atoms with the same number of protons but different numbers of neutrons

relative atomic mass - the average mass number for an element

## Equations, mixtures and chromatography

### chemical formulas and equations

chemical formula - shows the proportion of atoms of each element in a compound

chemical equation - shows the overall change in a reaction

example:

reactants

products

word equation: methane + oxygen → carbon dioxide + water

symbol equation:  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

chromatogram

insoluble components stay on the base line



submerge it

components separate as they move up the paper

## nuclear symbols

mass number = total number of protons and neutrons in an atom



atomic number = number of protons in an atom

## compounds

compound - a substance formed from two or more elements chemically bonded together in fixed proportions

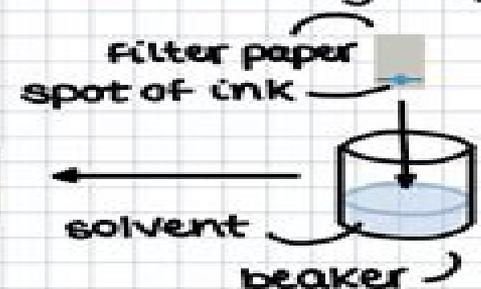
molecule - particle containing two or more non-metal atoms bonded covalently. Molecules can be elements or compounds

## mixtures

mixtures - substances made up of different elements or compounds that are not chemically bonded to each other

mixtures can be separated using physical methods

## paper chromatography



# Chapter 7 Notes Atomic Structure And Periodicity

**Arthur L. Williams, Harland D.  
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