

# 9

## Thermodynamics

1. Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
- (a)  $\Delta S$  is negative and therefore,  $\Delta H$  should be highly positive  
(b)  $\Delta S$  is negative and therefore,  $\Delta H$  should be highly negative  
(c)  $\Delta S$  is positive and therefore,  $\Delta H$  should be negative  
(d)  $\Delta S$  is positive and therefore,  $\Delta H$  should also be highly positive
2. For the reaction,  $X_2O_4(l) \longrightarrow 2XO_2(g)$ ,  $\Delta U = 2.1 \text{ kcal}$ ,  $\Delta S = 20 \text{ cal K}^{-1}$  at 300 K. Hence,  $\Delta G$  is [2014]
- (a) 2.7 kcal (b) - 2.7 kcal  
(c) 9.3 kcal (d) - 9.3 kcal
3. For a given exothermic reaction,  $K_p$  and  $K_p'$  are the equilibrium constants at temperatures  $T_1$  and  $T_2$ , respectively. Assuming that heat of reaction is constant in temperature range between  $T_1$  and  $T_2$ , it is readily observed that [2014]
- (a)  $K_p > K_p'$   
(b)  $K_p < K_p'$   
(c)  $K_p = K_p'$   
(d)  $K_p = \frac{1}{K_p'}$
4. A reaction having equal energies of activation for forward and reverse reactions has [2013]
- (a)  $\Delta S = 0$  (b)  $\Delta G = 0$   
(c)  $\Delta H = 0$  (d)  $\Delta H = \Delta G = \Delta S = 0$
5. In which of the following reactions, standard reaction entropy changes ( $\Delta S^\circ$ ) is positive and standard Gibbs energy change ( $\Delta G^\circ$ ) decreases sharply with increasing temperature? [2012]
- (a)  $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$   
(b)  $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$   
(c)  $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$   
(d)  $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
6. The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
- (a) 10.52 cal/mol K (b) 21.04 cal/mol K  
(c) 5.260 cal/mol K (d) 0.526 cal/mol K
7. Standard enthalpy of vaporisation  $\Delta_{\text{vap}}H^\circ$  for water at 100°C is 40.66 kJ mol<sup>-1</sup>. The internal energy of vaporisation of water at 100°C (in kJ mol<sup>-1</sup>) is (assume water vapour to behave like an ideal gas). [2012]
- (a) + 37.56 (b) - 43.76  
(c) + 43.76 (d) + 40.66
8. If the enthalpy change for the transition of liquid water to steam is 30 kJ mol<sup>-1</sup> at 27°C, the entropy change for the process would be [2011]
- (a) 1.0 J mol<sup>-1</sup>K<sup>-1</sup>  
(b) 0.1 J mol<sup>-1</sup>K<sup>-1</sup>  
(c) 100 J mol<sup>-1</sup>K<sup>-1</sup>  
(d) 10 J mol<sup>-1</sup>K<sup>-1</sup>

# Solution Thermodynamics Important Questions And Answers

**John H. Meiser**



## **Solution Thermodynamics Important Questions And Answers:**

**Comprehensive Physics for AIEEE** Prof Narinder Kumar, 2019      Physical Chemistry of Polymer Solutions K.

Kamide, T. Dobashi, 2000-10-16 This book is mainly concerned with building a narrow but secure ladder which polymer chemists or engineers can climb from the primary level to an advanced level without great difficulty but by no means easily either This book describes some fundamentally important topics carefully chosen covering subjects from thermodynamics to molecular weight and its distribution effects For help in self education the book adopts a Questions and Answers format The mathematical derivation of each equation is shown in detail For further reading some original references are also given Numerous physical properties of polymer solutions are known to be significantly different from those of low molecular weight solutions The most probable explanation of this obvious discrepancy is the large molar volume ratio of solute to solvent together with the large number of consecutive segments that constitute each single molecule of the polymer chains present as solute Thorough understanding of the physical chemistry of polymer solutions requires some prior mathematical background in its students In the original literature detailed mathematical derivations of the equations are universally omitted for the sake of space saving and simplicity In textbooks of polymer science only extremely rough schemes of the theories and then the final equations are shown As a consequence the student cannot learn unaided the details of the theory in which he or she is interested from the existing textbooks however without a full understanding of the theory one cannot analyze actual experimental data to obtain more basic and realistic physical quantities In particular if one intends to apply the theories in industry accurate understanding and ability to modify the theory are essential      **Fundamental Solutions**

**for Differential Operators and Applications** Prem Kythe, 2012-12-06 Overview Many problems in mathematical physics and applied mathematics can be reduced to boundary value problems for differential and in some cases integrodifferential equations These equations are solved by using methods from the theory of ordinary and partial differential equations variational calculus operational calculus function theory functional analysis probability theory numerical analysis and computational techniques Mathematical models of quantum physics require new areas such as generalized functions theory of distributions functions of several complex variables and topological and algebraic methods The main purpose of this book is to provide a self contained and systematic introduction to just one aspect of analysis which deals with the theory of fundamental solutions for differential operators and their applications to boundary value problems of mathematical physics applied mathematics and engineering with the related applicable and computational features The subject matter of this book has its own deep rooted theoretical importance since it is related to Green's functions which are associated with most boundary value problems The application of fundamental solutions to a recently developed area of boundary element methods has provided a distinct advantage in that an integral equation representation of a boundary value problem is often x PREFACE more easily solved by numerical methods than a differential equation with specified boundary and initial conditions

This situation makes the subject more attractive to those whose interest is primarily in numerical methods

*Comprehensive Objective Physics* Narinder Kumar, Dr. J. K. Juneja, 2006      **Selected Solutions for Chemistry, Concepts and Models by Robinson, Odom, and Holtzclaw** John H. Meiser, 1992      **Solid Solutions in Silicate and Oxide systems** C.A. Geiger, 2001-01-01      **Radiation Processes in Crystal Solid Solutions** Gennadi Gladyshev, 2012-03-31

Radiation Processes In Crystal Solid Solutions is a monograph explaining processes occurring in two classes of crystal solids metal alloys and doped alkali halides under irradiation by various types of radiation alpha beta gamma X radiations ions While metal alloys may differ in high radiation stability solid solutions based on alkali halides are very radiation sensitive materials Radiation defect production mechanisms intrinsic and extrinsic radiation defects a role of complexes an impurity radiation defect which explain distinction in radiating stability of the specified classes of solid solutions are discussed in this e book To describe radiation induced phase transformations two approaches are highlighted kinetic and thermodynamic This e book also includes research on the effect of small radiation doses in a structurally solid phase state of a solution along with a semi quantitative estimation of radiation effects with respect to temperature changes This e book should be a useful reference for advanced readers interested in the physics of radiation and solid state physics

*Proceedings of the Royal Society of London* Royal Society (Great Britain), 1912 Publishes research papers in the mathematical and physical sciences Continued by *Proceedings Mathematical and physical sciences* and *Proceedings Mathematical physical and engineering sciences*      *Industrial & Engineering Chemistry* , 1925      **Nature** Sir Norman Lockyer, 1906      *Journal of the American Chemical Society* American Chemical Society, 1923 Proceedings of the Society are included in v 1 59 1879 1937      **The General Principles** Daniel Frost Comstock, 1911      **Physical Chemistry : Solutions Manual** Robert A. Alberty, 1980      **The Technology Monthly and Harvard Engineering Journal** , 1916      **Technology Monthly and Harvard Engineering Journal** , 1916      **Osmotic Pressure** Faraday Society, 1917      *Transactions of the Faraday Society* Faraday Society, 1918      *Scientific American* , 1891      *Russian Journal of Physical Chemistry* , 2001

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