

Solutions Manual to Accompany

# Basic Principles and Calculations in Chemical Engineering

Seventh Edition

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PRENTICE HALL  
Professional Technical Reference  
Upper Saddle River, New Jersey 07458

**TYPICAL EXAMS FOR A ONE SEMESTER COURSE**  
(Scheduled in the evening to avoid time constraints on students)

**Exam No. 1**  
(Open Book, 1 1/2 hours)

**PROBLEM 1 (25%)**

Hydrogen can be separated from natural gas by diffusion through a metal tube. The rate of separation is given by

$$N = 2 \pi DpR$$

where

- N = rate of transport of  $H_2$  from the tube, g moles/decade (of length of tube)
- D = diffusion coefficient
- p = molar density of  $H_2$ , g moles/cm<sup>3</sup>
- R = log mass radius of tube,  $r_2 - r_1 / \ln \left( \frac{r_2}{r_1} \right)$ , with r in cm.

What are the units of D?

**PROBLEM 2 (25%)**

A pellet of bones weighing 10 tons is dropped from a lift truck from a height of 10 feet. The maximum velocity that pellet attains before hitting the ground is 8 ft/sec. How much kinetic energy does the pellet have in (Joules) at this velocity?

**PROBLEM 3 (25%)**

The specific gravity of a fuel oil is 0.82. What is the density of the oil in lb/ft<sup>3</sup>? Show all units.

**PROBLEM 4 (25%)**

Sulfur trioxide ( $SO_3$ ) can be absorbed in sulfuric acid solution to form more concentrated sulfuric acid. If the gas to be absorbed contains 50%  $SO_3$ , 41%  $N_2$ , 3%  $SO_2$ , and 1%  $O_2$ , how many parts per million of  $O_2$  are there in the gas? (a) What is the composition of the gas on a  $H_2$  free basis?

**PROBLEM 5 (25%)**

You have 100 kilograms of gas of the following composition:

$CH_4$	30%
$H_2$	20%
$N_2$	50%

What is the average molecular weight of this gas?

**PROBLEM 6 (25%)**

If the heat capacity of a substance is 3.32 J/g(°C) and its molecular weight is 75.4, what is its heat capacity in

- (a) Btu(°F)
- (b) J/kg(°K)
- (c) J/gmol(°C)

**PROBLEM 7 (25%)**

A rock containing 100%  $BaSO_4$  is heated with coke (94% C, 6% ash) and the composition of the product is  $BaSO_4$  (11.5%),  $BaS$  (72.9%), C(13.9%), ash(2.2%). The reaction is



Calculate the percent excess reactant, and the degree of completion of the reaction.

**PROBLEM 8 (25%)**

A gas cylinder to which is attached an Bourdon gage appears to be at a pressure of 21.38 in.  $H_2O$  at 70°F. The barometer reads 141.8 inHg. A student claims that the pressure in the tank is 1.3 psia, but another student points out that this is impossible - the pressure is really 28.2 psia. Can 1.3 psia be correct? Explain and show calculations to back up your explanation.

# Solution Basic Principles Himmelblau Ed 6

**Jin-Ying Zhang**



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*Chemical Process Safety* Daniel A. Crowl, Joseph F. Louvar, 2001-10-16 Combines academic theory with practical industry experience Updated to include the latest regulations and references Covers hazard identification risk assessment and inherent safety Case studies and problem sets enhance learning Long awaited revision of the industry best seller This fully revised second edition of *Chemical Process Safety Fundamentals with Applications* combines rigorous academic methods with real life industrial experience to create a unique resource for students and professionals alike The primary focus on technical fundamentals of chemical process safety provides a solid groundwork for understanding with full coverage of both prevention and mitigation measures Subjects include Toxicology and industrial hygiene Vapor and liquid releases and dispersion modeling Flammability characterization Relief and explosion venting In addition to an overview of government regulations the book introduces the resources of the AIChE Center for Chemical Process Safety library Guidelines are offered for hazard identification and risk assessment The book concludes with case histories drawn directly from the authors experience in the field A perfect reference for industry professionals *Chemical Process Safety Fundamentals with Applications* Second Edition is also ideal for teaching at the graduate and senior undergraduate levels Each chapter includes 30 problems and a solutions manual is now available for instructors

*Chemical Engineering Design and Analysis* T. Michael Duncan, Jeffrey A. Reimer, 2019-01-24 The go to guide to learn the principles and practices of design and analysis in chemical engineering

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Colin William Oloman, 2023-06-20 Material and energy M E balances are fundamental to biological chemical electrochemical photochemical and environmental engineering disciplines and important in many fields related to sustainable development This comprehensive compendium presents the basic M E balance concepts and calculations in a format easily digested by students engineering professionals and those concerned with related environmental issues The useful reference text includes worked examples for each chapter and demonstrates process balances in the framework of M E concerns of the 21st century The additional problems and solutions in the Appendix embrace a wide range of subjects from fossil fuels to fuel cells solar energy space stations carbon dioxide capture and sodium ion batteries

*Books in Print*, 1981 **AIChEMI Modular Instruction** American Institute of Chemical Engineers, 1981 *Books in Print Supplement*, 1985 *Solution Manual to Accompany Basic Principles and Calculations in Chemical Engineering* David Mautner Himmelblau, James B. Riggs, 2004-01

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### **Table of Contents Solution Basic Principles Himmelblau Ed 6**

1. Understanding the eBook Solution Basic Principles Himmelblau Ed 6
  - The Rise of Digital Reading Solution Basic Principles Himmelblau Ed 6
  - Advantages of eBooks Over Traditional Books
2. Identifying Solution Basic Principles Himmelblau Ed 6
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Solution Basic Principles Himmelblau Ed 6
  - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Basic Principles Himmelblau Ed 6
  - Personalized Recommendations

- Solution Basic Principles Himmelblau Ed 6 User Reviews and Ratings
- Solution Basic Principles Himmelblau Ed 6 and Bestseller Lists
- 5. Accessing Solution Basic Principles Himmelblau Ed 6 Free and Paid eBooks
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  - ePub, PDF, MOBI, and More
  - Solution Basic Principles Himmelblau Ed 6 Compatibility with Devices
  - Solution Basic Principles Himmelblau Ed 6 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Solution Basic Principles Himmelblau Ed 6
  - Highlighting and Note-Taking Solution Basic Principles Himmelblau Ed 6
  - Interactive Elements Solution Basic Principles Himmelblau Ed 6
- 8. Staying Engaged with Solution Basic Principles Himmelblau Ed 6
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Solution Basic Principles Himmelblau Ed 6
- 9. Balancing eBooks and Physical Books Solution Basic Principles Himmelblau Ed 6
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Solution Basic Principles Himmelblau Ed 6
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Solution Basic Principles Himmelblau Ed 6
  - Setting Reading Goals Solution Basic Principles Himmelblau Ed 6
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solution Basic Principles Himmelblau Ed 6
  - Fact-Checking eBook Content of Solution Basic Principles Himmelblau Ed 6

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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