

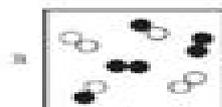
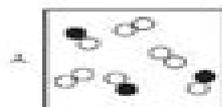
8. Given the equation representing a system at equilibrium:



Which statement describes this system?

- A. The concentration of $\text{PCl}_5(\text{g})$ is increasing.
- B. The concentration of $\text{PCl}_3(\text{g})$ is decreasing.
- C. The concentrations of $\text{PCl}_3(\text{g})$ and $\text{PCl}_5(\text{g})$ are equal.
- D. The concentrations of $\text{PCl}_3(\text{g})$ and $\text{PCl}_5(\text{g})$ are constant.

9. Which particle diagram represents a mixture of three substances?



10. Which change in the H^+ ion concentration of an aqueous solution represents a decrease of one unit on the pH scale?

- A. a tenfold increase
- B. a tenfold decrease
- C. a hundredfold increase
- D. a hundredfold decrease

11. How many joules of heat are absorbed to raise the temperature of 435 grams of water at 1 atm from 25°C to its boiling point, 100°C ?

- A. $4.5 \times 10^4 \text{ J}$
- B. $1.4 \times 10^5 \text{ J}$
- C. $2.5 \times 10^7 \text{ J}$
- D. $7.4 \times 10^7 \text{ J}$

12. Based on Table F, which compound is *least* soluble in water?

- A. AlPO_4
- B. Li_2SO_4
- C. $\text{Ca}(\text{OH})_2$
- D. $\text{AgC}_2\text{H}_3\text{O}_2$



Which statement describes the energy changes that occur in this reaction?

- A. Energy is absorbed as bonds are formed, only.
- B. Energy is released as bonds are broken, only.
- C. Energy is absorbed as bonds are formed, and energy is released as bonds are broken.
- D. Energy is absorbed as bonds are broken, and energy is released as bonds are formed.

14. What is the gram-formula mass of $\text{Ca}(\text{OH})_2$?

- A. 29 g/mol
- B. 54 g/mol
- C. 57 g/mol
- D. 74 g/mol

15. Given the balanced equation representing a reaction:



Which mass of oxygen completely reacts with 4.0 grams of hydrogen to produce 36.0 grams of water?

- A. 8.0 g
- B. 16.0 g
- C. 32.0 g
- D. 40.0 g

16. Which molecular formula is also an empirical formula?

- A. C_6H_6
- B. H_2O_2
- C. N_2H_4
- D. N_2O_5

17. What is the chemical name for Na_2SO_3 ?

- A. sodium sulfite
- B. sodium sulfate
- C. sodium sulfide
- D. sodium thiosulfate

18. Which general trends in atomic radius and electronegativity are observed as the elements in Period 3 are considered in order of increasing atomic number?

- A. Atomic radius decreases and electronegativity increases.
- B. Atomic radius increases and electronegativity decreases.
- C. Both atomic radius and electronegativity increase.
- D. Both atomic radius and electronegativity decrease.

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