

SALT ANALYSIS

SCHEME OF ANALYSIS

EXPERIMENT	OBSERVATION	INFERENCE
PRELIMINARY EXAMINATIONS :		
1) Noted the colour	Colourless	Absence of $\text{Cu}^{2+}, \text{Fe}^{2+}, \text{Fe}^{3+}, \text{Mn}^{2+}, \text{Co}^{2+}, \text{Ni}^{2+}$
	Blue	Pre. of Cu^{2+}
	Brown	Pre. of Fe^{3+}
	Green	Pre. of Ni^{2+}
	Pale pink	Pre. of Mn^{2+}
2) Noted the solubility a) In water b) In dil.HCl	a) Soluble and no ppt with Na_2CO_3 b) Insoluble	a) Pre. of NH_4^+ b) Pre. of Pb^{2+}
3) Flame Test : A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame.	a) Blue surrounded by green b) Brick red c) Crimson red d) Bluish white e) Grassy green on prolonged heating	a) Pre. of Cu^{2+} b) Pre. of Ca^{2+} c) Pre. of Sr^{2+} d) Pre. of Pb^{2+} e) Pre. of Ba^{2+}
4) Ash Test : A little of the salt is heated with con. HNO_3 and cobalt nitrate solution. A filter paper is dipped in it and burnt.	a) Green ash b) Blue tinted ash c) Pale pink ash	a) Pre. of Zn^{2+} b) Pre. of Al^{3+} c) Pre. of Mg^{2+}
SYSTEMATIC ANALYSIS OF ANIONS :		
1) Dil. H_2SO_4 test : To the salt add dil. H_2SO_4 , warm if required.	a) Evolution of Colourless, odourless gas with brisk effervescence, turns limewater milky. b) Smell of vinegar	a) Pre. of carbonates b) Pre. of acetate

Chemistry Practicals For Class 11 Salt Analysis

K Morrison



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Table of Contents Chemistry Practicals For Class 11 Salt Analysis

1. Understanding the eBook Chemistry Practicals For Class 11 Salt Analysis
 - The Rise of Digital Reading Chemistry Practicals For Class 11 Salt Analysis
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practicals For Class 11 Salt Analysis
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Practicals For Class 11 Salt Analysis
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chemistry Practicals For Class 11 Salt Analysis
 - Personalized Recommendations
 - Chemistry Practicals For Class 11 Salt Analysis User Reviews and Ratings
 - Chemistry Practicals For Class 11 Salt Analysis and Bestseller Lists
5. Accessing Chemistry Practicals For Class 11 Salt Analysis Free and Paid eBooks
 - Chemistry Practicals For Class 11 Salt Analysis Public Domain eBooks
 - Chemistry Practicals For Class 11 Salt Analysis eBook Subscription Services
 - Chemistry Practicals For Class 11 Salt Analysis Budget-Friendly Options
6. Navigating Chemistry Practicals For Class 11 Salt Analysis eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Practicals For Class 11 Salt Analysis Compatibility with Devices
 - Chemistry Practicals For Class 11 Salt Analysis Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Practicals For Class 11 Salt Analysis
 - Highlighting and Note-Taking Chemistry Practicals For Class 11 Salt Analysis
 - Interactive Elements Chemistry Practicals For Class 11 Salt Analysis
8. Staying Engaged with Chemistry Practicals For Class 11 Salt Analysis
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practicals For Class 11 Salt Analysis
9. Balancing eBooks and Physical Books Chemistry Practicals For Class 11 Salt Analysis
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practicals For Class 11 Salt Analysis
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Practicals For Class 11 Salt Analysis
 - Setting Reading Goals Chemistry Practicals For Class 11 Salt Analysis
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Practicals For Class 11 Salt Analysis
 - Fact-Checking eBook Content of Chemistry Practicals For Class 11 Salt Analysis
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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