

Chemical Reactions and Equations

Chemical Reaction

- change in state
- colour
- gas
- temperature

ex: Magnesium + Oxygen \rightarrow

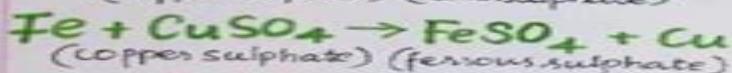
(Reactants)

magnesium oxide
(products)



DISPLACEMENT

Reaction

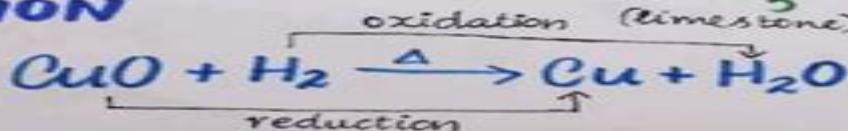


DOUBLE DISPLACEMENT

Reaction



OXIDATION & REDUCTION



Why balance Chemical equation

\Rightarrow Mass can neither be created nor destroyed in a chemical reaction.
 \triangleright Total mass of elements in the product of a chemical reaction has to be equal to the total mass of the elements in the reactants side.

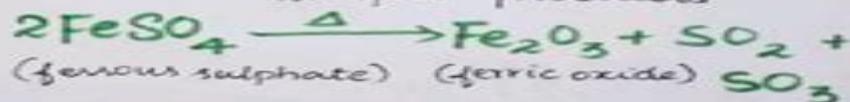
COMBINATION

Reaction when two or more substances combine to form a single product.



DECOMPOSITION

Reaction when a single reactant breaks down to simpler products



Chemistry Notes Chapter 8 Chemical Reactions

Camilla Rothe



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Basic Chemical Thermodynamics Eric Brian Smith,1982 The major change in this edition is the addition of a new chapter which provides a simple treatment of the molecular basis of thermodynamics Numerical examples have been added but overall emphasis is again placed on a clear explanation of the underlying concepts of thermodynamics the traditional approaches to the subject are modified so that the reader is led from an examination of equilibrium in the everyday world of mechanical objects to a thorough understanding of the factors that determine equilibrium in chemical systems The results are applied to a wide range of physicochemical problems The book also gives a brief introduction to the difficult problems of non ideal solutions and the concept of activity Meteorologische Zeitschrift ,2001 The Chemist and Druggist ,1921

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