

CHEMISTRY MODULES OBJECTIVE AND SYLLABUS

LEVEL 1 MODULES

CM1111: Inorganic Chemistry 1 (Essential Module)

Workload: 4-1-0-2-3

Aims & Objectives

This module introduces some basic concepts of modern inorganic chemistry. In the first part, the atomic structure, the bonding models (Lewis structure, valence bond theory, basic molecular theory), basic concepts of acids and bases, redox reactions will be covered. In the part 2, the periodicity and chemistry of most main group elements will be included.

Syllabus

Part One: Basic Concepts

1. Atomic Structure
 - 1) Atomic Quantization
 - 2) Schrödinger Equation for hydrogen atom
 - 3) Many-electron systems
2. Chemical periodicity and the periodic Table
3. Molecular structure and bonding
 - 1) Lewis structure
 - 2) Valence bond theory
 - 3) Basic molecular orbital theory
4. Acid - Base and donor - acceptor chemistry
 - 1) Brønsted-Lowry concept
 - 2) Lewis concept
5. Reduction and Oxidation

Part Two: Chemistry of Selected Main Group Elements

1. Hydrogen and Group 1
2. Group 2
3. Group 13
4. Group 14
5. Group 15
6. Group 16
7. Group 17
8. Group 18

CM1121: Organic Chemistry 1 (Essential Module)

Workload: 4-2-0-2-2

Aims & Objectives

The module deals primarily with the basic principles to understand the structure and reactivity of organic molecules. Emphasis is on substitution and elimination reactions and chemistry of various functional groups. You will be taught the basic concepts on how simple molecules can be constructed. Reactions mechanism, organic transformations and stereochemistry will also be discussed.

By reading this modules, students should be able to:

- i. Understand and be able to apply and evaluate simple organic reaction transformations, functional group interconversion and C-C bond formation reactions.
- ii. Understand the scope and limitations as well as the mechanisms of organic reactions.
- iii. Understand the importance of stereochemistry and be able to view three-dimensional structures with ease.
- iv. Describe and explain currently held views of chemical bonding, and account for the chemical reactivity of reactions.

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RJ Shavelson



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