

Empirical and Molecular Formulas Worksheet

Key

Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l}
 \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} \approx 1.96 \quad 2 \\
 \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} \approx 5.88 \quad 6 \\
 \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1
 \end{array}$$

$\text{C}_2\text{H}_6\text{O}$

2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{array}{l}
 \text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5} \\
 \text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5
 \end{array}$$

Au_2O_3

3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\begin{array}{l}
 \text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad \cdot \quad 2 \\
 \text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad \cdot \quad 8 \\
 \text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad \cdot \quad 1 \\
 \text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad \cdot \quad 4
 \end{array}$$

$\text{N}_2\text{H}_8\text{S}_4\text{O}_4$

Molecular Formula

4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) = 30$$

$$\frac{180}{30} = 6 \quad \text{C}_6\text{H}_{12}\text{O}_6$$

5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\begin{array}{l}
 \text{C} \quad \frac{1.388}{3.585} \rightarrow 38.7\% \rightarrow \frac{3.87\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1 \\
 \text{H} \quad \frac{0.345}{3.585} \rightarrow \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3 \\
 \text{O} \quad \frac{1.850}{3.585} \rightarrow \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1
 \end{array}$$

$\text{C}_2\text{H}_6\text{O}_2$

$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$

$$\frac{62}{31.03} = 2$$

1	$\text{C}_2\text{H}_6\text{O}$
2	Au_2O_3
3	NH_4SO_4
4	$\text{C}_2\text{H}_6\text{O}_2$
5	$\text{C}_2\text{H}_6\text{O}$

Chemistry Empirical Formula Worksheet Answers

Katrin Zwirgmaier



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Table of Contents Chemistry Empirical Formula Worksheet Answers

1. Understanding the eBook Chemistry Empirical Formula Worksheet Answers
 - The Rise of Digital Reading Chemistry Empirical Formula Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Empirical Formula Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Empirical Formula Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Empirical Formula Worksheet Answers
 - Personalized Recommendations
 - Chemistry Empirical Formula Worksheet Answers User Reviews and Ratings
 - Chemistry Empirical Formula Worksheet Answers and Bestseller Lists
5. Accessing Chemistry Empirical Formula Worksheet Answers Free and Paid eBooks
 - Chemistry Empirical Formula Worksheet Answers Public Domain eBooks
 - Chemistry Empirical Formula Worksheet Answers eBook Subscription Services
 - Chemistry Empirical Formula Worksheet Answers Budget-Friendly Options
6. Navigating Chemistry Empirical Formula Worksheet Answers eBook Formats

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 - Chemistry Empirical Formula Worksheet Answers Compatibility with Devices
 - Chemistry Empirical Formula Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Empirical Formula Worksheet Answers
 - Highlighting and Note-Taking Chemistry Empirical Formula Worksheet Answers
 - Interactive Elements Chemistry Empirical Formula Worksheet Answers
 8. Staying Engaged with Chemistry Empirical Formula Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Empirical Formula Worksheet Answers
 9. Balancing eBooks and Physical Books Chemistry Empirical Formula Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Empirical Formula Worksheet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Empirical Formula Worksheet Answers
 - Setting Reading Goals Chemistry Empirical Formula Worksheet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Empirical Formula Worksheet Answers
 - Fact-Checking eBook Content of Chemistry Empirical Formula Worksheet Answers
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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