

Julianne  
Partner: Caden Long  
Mrs. Reid (Chemistry)  
September 28, 2021

**Title:** What Is Chemical Reaction?

**Introduction:** The goal of this experiment is to determine chemical properties of hydrochloric acid and copper chloride; identify the evidence that we can gather to find out which chemical change that occurred, whether in nature or laboratory; if the law of conservation of mass applies to a sample chemical reaction in one of these series. A chemical change is defined as a variable composition and properties of substance. There are many types of observable changes such as formation of solid precipitate upon mixing two solutions; release of gas bubbles that are not due to a physical change; a color change that does not result from dilution or color mixing; or a temperature change that is not caused by the external heating or cooling. For instance, a nail can become tarnished when it gets old as new substances begin to take the place of the old substances of the nail. Still, it is physically a nail but the "formula" differentiates. The nail now has been oxidized. If you are a cook, you are now a food scientist. Chemical changes don't have to occur in the lab all the time, they can actually be right there in your kitchen. Think of a simple breakfast recipe like scrambled eggs. Cooking an egg uses a chemical reaction because the heat polymerizes proteins in egg white, while the hydrogen and sulfur in the yolk can react to form hydrogen sulfide gas. Same thing happens with cooked meats, and cake.

**Objective:** to examine the chemical properties of hydrochloric acid and copper chloride, to identify the types of reactions that they undergo, and to determine the law of conservation of mass applies to a sample chemical reaction.

**Hypothesis:** as acid is a hazardous chemical, the samples might experience corroding or discolor as soaking in the hydrochloric solution. The samples may not react as dramatically to copper chloride as hydrochloric because it does not have an acid name.

**Result:**

Data Table A: Reactions of Hydrochloric Acid

Initial appearance of HCl Solutions: transparent, and clear liquid

Initial temperature of HCl solutions: 74° F

Reaction Well	Reagents	Observations
1	HCl + NaOH	After 1 minute: 72° F Litmus paper turns pink.
2	HCl + NaHCO <sub>3</sub>	Bubble frizzing, temperature decreases
3	HCl + AgNO <sub>3</sub>	Foam occurs in the beginning, then the solution

# Chemical Reaction Lab Report Grade 10

**ML Morrison**



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## **Table of Contents Chemical Reaction Lab Report Grade 10**

1. Understanding the eBook Chemical Reaction Lab Report Grade 10
  - The Rise of Digital Reading Chemical Reaction Lab Report Grade 10
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Reaction Lab Report Grade 10
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Reaction Lab Report Grade 10
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Reaction Lab Report Grade 10
  - Personalized Recommendations
  - Chemical Reaction Lab Report Grade 10 User Reviews and Ratings
  - Chemical Reaction Lab Report Grade 10 and Bestseller Lists

5. Accessing Chemical Reaction Lab Report Grade 10 Free and Paid eBooks
  - Chemical Reaction Lab Report Grade 10 Public Domain eBooks
  - Chemical Reaction Lab Report Grade 10 eBook Subscription Services
  - Chemical Reaction Lab Report Grade 10 Budget-Friendly Options
6. Navigating Chemical Reaction Lab Report Grade 10 eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Reaction Lab Report Grade 10 Compatibility with Devices
  - Chemical Reaction Lab Report Grade 10 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Reaction Lab Report Grade 10
  - Highlighting and Note-Taking Chemical Reaction Lab Report Grade 10
  - Interactive Elements Chemical Reaction Lab Report Grade 10
8. Staying Engaged with Chemical Reaction Lab Report Grade 10
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Reaction Lab Report Grade 10
9. Balancing eBooks and Physical Books Chemical Reaction Lab Report Grade 10
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Reaction Lab Report Grade 10
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemical Reaction Lab Report Grade 10
  - Setting Reading Goals Chemical Reaction Lab Report Grade 10
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Reaction Lab Report Grade 10
  - Fact-Checking eBook Content of Chemical Reaction Lab Report Grade 10
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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