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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$ H_2O L $2\text{Cr}^{3+} \text{aq}$ $2\text{OH}^- \text{aq}$

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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq}$ $2\text{PO}_4^{3-} \text{aq}$ $\text{Ba}_3\text{PO}_4 \text{s}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water h aq oh aq h₂ o l the spectator ions in koh aq and hno₃ aq are k aq and no₃ aq how to identify and list the

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 $2\text{Cl}^{-}\text{aq} + \text{Cu}^{2+}\text{aq} \rightarrow \text{CuCl}_2\text{s}$
 $2\text{Cl}^{-}\text{aq} + \text{Cu}^{2+}\text{aq} \rightarrow \text{CuCl}_2\text{s}$ these spectator ions are not actively involved in the reaction the existence of
spectator ions affects the debye length