

Introduction to Strong and Weak Acids and Bases PhET Lab (revd 5/2011)*How does the strength of an acid or base affect conductivity/pH?***Introduction:**

When you test your pool's pH, what are those little vials or paper strips telling you? When you hear an acid called "strong" or "weak", what do those terms refer to? In aqueous solutions, compounds can exist as molecules (*undissociated*) or ions (*dissociated*). When an acid or a base exists in solution nearly completely as dissociated ions, we refer to that acid or base as *strong*. A *weak* acid or base will donate ions to the solution, but will remain primarily as undissociated molecules.

**Acid-Base Solutions****Notation:**

Acids are abbreviated HA, with the H representing the proton (H⁺) the acid donates to the solution. The A is referred to as the acidic anion (A⁻) that is left in solution as the proton is donated. $HA \rightleftharpoons H^+ + A^-$

Strong Bases are abbreviated MOH, with the OH representing the hydroxide ion (OH⁻) the base donates to the solution. The M is cation (M⁺) that is left in solution as the hydroxide is donated. $MOH \rightarrow M^+ + OH^-$.

Autoionization:

Even without any acid or base added a very small number of water molecules will form protons (H⁺) and hydroxide ions (OH⁻). The protons will then form **hydronium ions**, the acid ion.



Procedure: *PhET Simulations* → *Play With Sims* → *Chemistry* → *Acid-Base Solutions* → [Run Now!](#)

The concentration of the acids and bases used in the [Introduction](#) at 0.010 (10⁻²) Molar.

- Begin with a **strong acid** and lower the pH probe into the beaker. What is the pH of this solution?
- Test this strong acid with both pH paper and the conductivity probe. What color does the pH indicator become? Is this strong acid an electrolyte? Does current travel through this solution?
- Repeat the above tests with the weak acid, the strong base, and the weak base, and water. Collect your observations in the table below:

	Strong Acid	Weak Acid	Strong Base	Weak Base	Water
pH meter read (value)	2.00	4.50	12.00	9.50	7.00
pH paper (color)	2 red	4 orange	11 blue	9 green	7 yellow
Conductivity (bright/dim/none)	Bright	dim	10 Right	dim	none
Exists as Mostly (ions/molecules)	ions	ions	molecules	molecules	ions

Acid And Bases Ph Phet Lab Answers

Adam Hulanicki



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Table of Contents Acid And Bases Ph Phet Lab Answers

1. Understanding the eBook Acid And Bases Ph Phet Lab Answers
 - The Rise of Digital Reading Acid And Bases Ph Phet Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid And Bases Ph Phet Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid And Bases Ph Phet Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid And Bases Ph Phet Lab Answers
 - Personalized Recommendations
 - Acid And Bases Ph Phet Lab Answers User Reviews and Ratings
 - Acid And Bases Ph Phet Lab Answers and Bestseller Lists
5. Accessing Acid And Bases Ph Phet Lab Answers Free and Paid eBooks
 - Acid And Bases Ph Phet Lab Answers Public Domain eBooks
 - Acid And Bases Ph Phet Lab Answers eBook Subscription Services
 - Acid And Bases Ph Phet Lab Answers Budget-Friendly Options

6. Navigating Acid And Bases Ph Phet Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Acid And Bases Ph Phet Lab Answers Compatibility with Devices
 - Acid And Bases Ph Phet Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid And Bases Ph Phet Lab Answers
 - Highlighting and Note-Taking Acid And Bases Ph Phet Lab Answers
 - Interactive Elements Acid And Bases Ph Phet Lab Answers
8. Staying Engaged with Acid And Bases Ph Phet Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid And Bases Ph Phet Lab Answers
9. Balancing eBooks and Physical Books Acid And Bases Ph Phet Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid And Bases Ph Phet Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acid And Bases Ph Phet Lab Answers
 - Setting Reading Goals Acid And Bases Ph Phet Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acid And Bases Ph Phet Lab Answers
 - Fact-Checking eBook Content of Acid And Bases Ph Phet Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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