

## *Material Balance Calculations*

(Incorporating calculations of ideal gas laws, unit operations and chemical reactions etc.)

### THEORETICAL BACKGROUND

Ideal Gas Law :

$$PV = nRT$$

where,  $P$  = Total pressure of system  
 $V$  = Volume  
 $n$  = Number of moles  
 $T$  = Absolute temperature  
 $R$  = Universal gas constant

Values of universal gas constant ( $R$ ).

In different units, the values of universal gas constant are as under :

<i>Unit</i>	<i>Numerical value of R</i>
$\frac{(\text{atm.}) (\text{lit.})}{\text{gm mole } ^\circ\text{K}}$	0.08205
$\frac{(\text{atm.}) \text{m}^3}{\text{kg. mole } ^\circ\text{K}}$	0.08205
$\frac{(\text{atm.}) (\text{cc})}{\text{gm mole } ^\circ\text{K}}$	82.05
$\frac{(\text{Psi}) \text{ft}^3}{\text{lb. mole } ^\circ\text{R}}$	10.73
$\frac{(\text{atm.}) \text{ft}^3}{\text{lb. mole } ^\circ\text{R}}$	0.73
$\frac{(\text{mm Hg}) \text{lit.}}{\text{gm mole } ^\circ\text{K}}$	62.359
$\frac{\text{kcal}}{\text{kg. mole } ^\circ\text{K}}$	1.987

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