



Physical Chemistry

1. If 2 moles of $\text{Mg}(\text{OH})_2$ reacts with 3 moles of H_3PO_4 . Find maximum moles of $\text{Mg}_3(\text{PO}_4)_2$ that will form?
(A) 2/3 (B) 3/2
(C) 4/9 (D) 1
2. If 1 M aqueous solution of NaOH is also 1 molal. Find density of solution?
(A) 1.1 g/mL (B) 1.04 g/mL
(C) 0.96 g/mL (D) 1.2 g/mL
3. Which of the given quantities will change if temperature is raised?
(A) Molality (B) weight fraction
(C) Mole fraction (D) Molarity
4. If radius of 3rd shell of He^+ is x then radius of 2nd shell of Li^{2+} is
(A) $\frac{8x}{27}$ (B) $\frac{27x}{8}$
(C) $\frac{9x}{4}$ (D) $\frac{4x}{9}$
5. Number of orbitals associated with $n = 4$ and $m_s = -1/2$ is?
(A) 16 (B) 8
(C) 10 (D) 14
6. Which set of quantum number represents the highest energy of an atom
(A) $n = 3, l = 1, m = 0, s = +1/2$
(B) $n = 3, l = 2, m = 2, s = -1/2$
(C) $n = 3, l = 0, m = 0, s = +1/2$
(D) $n = 4, l = 0, m = 0, s = -1/2$
7. Find correct statement about P vs $1/V$ curve?
(A) Non linear and decreasing
(B) Non linear and increasing
(C) Linear and decreasing
(D) Linear and increasing
8. Boyle's temperature for a given gas is
(A) $\frac{a}{Rb}$ (B) $\frac{2a}{Rb}$
(C) $\frac{8a}{27Rb}$ (D) $\frac{27a}{8Rb}$
9. Compressibility factor at Boyle's temperature and low pressure is given by
(A) $Z = 1 + \frac{Pb}{RT}$ (B) $Z = 1 - \frac{Pb}{RT}$
(C) $Z = 1 - \frac{aP}{RT}$ (D) $Z = 1$
10. If 3 moles of a diatomic gas is mixed with 2 moles of monoatomic gas then C_p of mixture of gas is
(A) 3.1 R (B) 3 R
(C) 2.9 R (D) 3.5 R
11. Enthalpy is equal to
(A) Isochoric work
(B) Isobaric work
(C) Adiabatic work
(D) Adiabatic work – isobaric work
12. In which of given cases entropy of system is constant.
(A) isothermal rev expansion
(B) Isobaric rev expansion.
(C) adiabatic rev expansion.
(D) free expansion.
13. Find enthalpy of formation of NH_3 gas if B_x $\text{N} = \text{N}$ is x and $\text{H} - \text{H}$ is Y and $\text{N} - \text{H}$ is Z
(A) $x + 3y - 3z$ (B) $\frac{x+3y-3z}{2}$
(C) $3z - x - 3y$ (D) $\frac{x+3z}{2} - 3z$
14. For which of given Reaction $\Delta H = \Delta U$
(A) $2\text{HI}(g) \rightleftharpoons \text{H}_2(g) + \text{I}_2(g)$
(B) $\text{CH}_4(g) + 2\text{O}_2(g) \longrightarrow \text{CO}_2(g) + 2\text{H}_2\text{O}(l)$
(C) $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$
(D) $\text{CO}_2(g) + \text{C}_{(\text{graphite})} \rightleftharpoons 2\text{CO}(g)$

Bsc Part First Physical Chemistry Question Paper

JR Anderson



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