

Buffer Solution

The diagram illustrates the effect of adding acid (HCl) to water and a buffer solution. It is divided into two columns. The left column shows a beaker of water (H_2O) with a pH meter reading 7.0. A red liquid is added to it, and the pH meter now reads 3.0. The right column shows a beaker of water plus a buffer ($H_2O + \text{Buffer}$) with a pH meter reading 7.0. A red liquid is added to it, and the pH meter now reads 6.9. A central box labeled "Adding acid (HCl)" has arrows pointing down to the transition between the two rows.

Left side: H_2O (pH 7.0) → Adding acid (HCl) → pH 3.0 (turns acidic)

Right side: $H_2O + \text{Buffer}$ (pH 7.0) → Adding acid (HCl) → pH 6.9 (Stable pH)

The diagram illustrates the effect of adding base (NaOH) to water and a buffer solution. It is divided into two columns. The left column shows a beaker of water (H_2O) with a pH meter reading 7.0. A blue liquid is added to it, and the pH meter now reads 11.0. The right column shows a beaker of water plus a buffer ($H_2O + \text{Buffer}$) with a pH meter reading 7.0. A blue liquid is added to it, and the pH meter now reads 7.1. A central box labeled "Adding base (NaOH)" has arrows pointing down to the transition between the two rows.

Left side: H_2O (pH 7.0) → Adding base (NaOH) → pH 11.0 (turns basic)

Right side: $H_2O + \text{Buffer}$ (pH 7.0) → Adding base (NaOH) → pH 7.1 (Stable pH)

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