

Chapter-4

Reaction Stoichiometry

- Stoichiometry is the study of the quantitative relationships between amounts of reactants used and amounts of products formed by a chemical reaction.
 - Chemical equations
 - The mole and molar mass
 - Chemical formulas
 - Mass relationships in equations
 - Limiting reactant

1

4.1. Interpreting Stoichiometric Coefficients

$4\text{Fe(s)} + 3\text{O}_2\text{(g)}$	\rightarrow	$2\text{Fe}_2\text{O}_3\text{(s)}$
Iron + oxygen	\rightarrow	Iron (III) oxide
4 atoms of Fe + 3 molecules of O_2	\rightarrow	2 formula units of Fe_2O_3
4 mol of Fe + 3 mol of O_2	\rightarrow	2 mol Fe_2O_3
223.4 g of Fe + 96 g of O_2	\rightarrow	319.4 g of Fe_2O_3
319.4 g reactants	\rightarrow	319.4 g products

- The chemical equation does not directly tell you anything about the masses of reactants and products. However, by converting the known mole quantities to mass.
 - Moles are converted to mass by multiplying their molar mass

$$4 \text{ mol Fe} \times \frac{55.85 \text{ g Fe}}{1 \text{ mol Fe}} = 223.4 \text{ g Fe}$$

Reaction Stoichiometry Lab Answers

**Harold R. Hunt, Toby F. Block, George
M. McKelvy**

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Research Based Undergraduate Science Teaching Dennis W. Sunal, Cynthia Szymanski Sunal, Emmett L. Wright, Cheryl L. Mason, Dean Zollman, 2014-07-08 RISE Volume 6 examines research theory and practice in teaching undergraduate science It focuses on developing science literacy and best practices for faculty professional development The volume includes research on reforming science courses and supporting underrepresented groups in STEM education

Stoichiometry Unit Project Luann Marie Decker, 1998

Experimental Chemistry Russell A. Geanangel, Wesley William Wendlandt, 1979-08

Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001

Chemistry: Molecules, Matter, and Change Media Activities Book Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher, 2000-01-15 The Media Activity Book MAB for Jones Atkins Chemistry molecules matter and change contains chapters with lists and descriptions of some of the media available as you study the chapter Each activity begin with a specific textbook reference Then you are given a time estimate of how long it will take to use the media An M media icon in the margin of the textbook means that media exists to support that area of text The media is found in three different places on the website and on two CDs

Laboratory Experiments for General Chemistry Harold R. Hunt, Toby F. Block, George M. McKelvy, 2002 This established manual focuses on using non hazardous materials to teach the experimental nature of general chemistry Experiments are written to address students of various academic backgrounds and differing interests and abilities in chemistry While most experiments can be conducted in a single three hour period some have been designed to be completed over an extended time to illustrate that chemical systems do not work at an arbitrary schedule Suggestions are provided for combining experiments of shorter length and similar pedagogy

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Using Multimedia Technology in Chemistry Pre-laboratory Preparation Jeffrey Glen Yoder, 2002

Improving Student Comprehension in Chemistry Laboratories Tracy Lynn Haroff, 2006

A Concrete Stoichiometry Unit for High School Chemistry Jennifer Louise Pakkala, 2006

Chemical Investigations Nancy Konigsberg Kerner, 1986

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